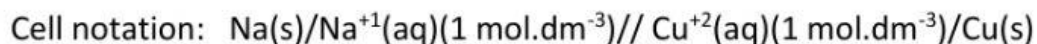
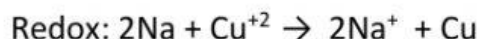
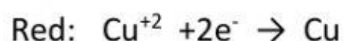


Eg 2)



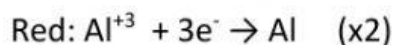
**notice the balancing number are not included in the cell notation*

Exercise:

Write the cell notation for the following:

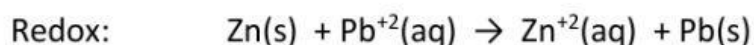
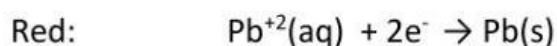
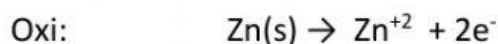
1)

Mg and Al



Cell notation:

2)



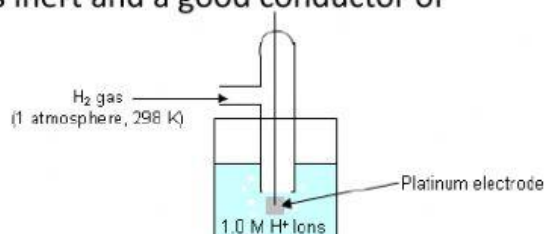
Cell notation:

Hydrogen half cell

Now you can imagine when any gas is involved it – you cannot make an electrode from a gas. Thus what happens is that you have a glass tube, in which you pump the gas into and the solution then has the same element in aqueous solution.

In the case of hydrogen gas- an acid is perfect as an electrolyte, since they all contain H^+ ions.

The electrode is a small Platinum electrode – since it is inert and a good conductor of electricity.



You'll also notice the value: 0V next to the hydrogen half reaction on the table of standard reduction potentials.

This is because Hydrogen is the **reference electrode**- meaning that no element by itself has a voltage, but if you connect 2 elements together that creates a potential difference. Hydrogen was thus used as the comparison to create the table.

Thus the standard H electrode is 0,00 V

The $E^\ominus$ is positive if electrons flow from the H to the other metal.

And negative if electrons flow from the other metal to Hydrogen.

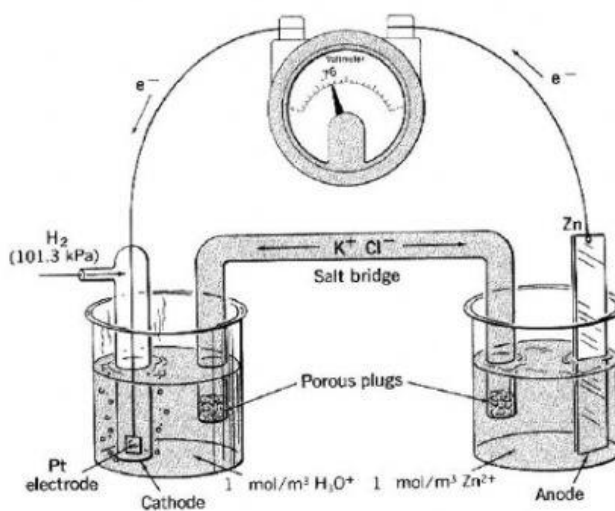
The standard-hydrogen half cell is used to compare other half cells to. It consists of a platinum electrode that is immersed in an acidic solution which contains $[H^+]$ of 1 mol.dm^{-3} . Hydrogen gas at a pressure of 1 atmosphere (101.3 kPa) is bubbled over the electrode. The temp is kept constant at 25°C .

Example

Oxidation half reaction: $\text{Zn} \rightarrow \text{Zn}^{+2} + 2e^-$

Reduction half reaction: $2H^+ + 2e^- \rightarrow H_2$

Redox/net reaction: $\text{Zn} + 2H^+ \rightarrow \text{Zn}^{+2} + H_2$



The only real difference is the cell notation – you need to include the Pt(s) in the hydrogen half of the notation.

Cell notation/Daniel cell notation:

$\text{Zn(s)} / \text{Zn}^{+2}(\text{aq})(1 \text{ mol.dm}^{-3}) // \text{H}^+(\text{aq})(1 \text{ mol.dm}^{-3}) / \text{H}_2(\text{g})(1 \text{ atm}) / \text{Pt(s)}$

Homework:

Pg 184 Q 4 (not 4.8)

Pg 185 Q5 (not 5.8)

Pg 186/187 Q6 and 7(not 6.8 or 7.8)

Pg 199 Q 9 (not 9.5)

Pg 201 Q10

Pg 203 Q8

Pg 204 Q8 (not 8.6)

Pg 208 Q10