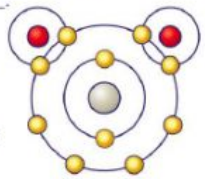
Chemical bonding

Name_____ Class_____ Date_____

1 A **covalent bond** usually forms **between** _____.

A two metals
B two metalloids
C a metal and a nonmetal
D two nonmetals

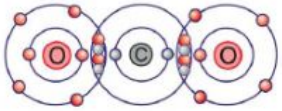


2 Using the **electron dot diagram** below, determine what kind of **bond** has formed between the carbon atom and the two oxygen atoms.

A metallic
B covalent
C ionic
D molecular

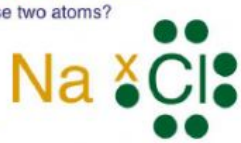
3 Using this diagram of carbon dioxide, determine how many **valence electrons** each **oxygen** atom has after bonding with the carbon atom.

A 8
B 6
C 4
D 2



4 In the diagram below, one **sodium** atom has bonded to one **chlorine** atom. What type of **bond** has formed between these two atoms?

A metallic
B covalent
C ionic
D molecular



5 This **metal** and **nonmetal** have formed an ionic bond. What is the name of the **compound** that has been formed?

A sodium chlorine
B sodium chloride
C sodiumide chloride
D sodiumide chloride

6 When a group of **active metal atoms** are put together, a strong force of **attraction** called a(n) _____ develops between these atoms.

A ionic bond
B covalent bond
C metallic bond
D magnetic bond

7 The **diagram** below **shows** that aluminum has bonded to chlorine. The proper term for this **diagram** is a _____.

A chemical formula
B compound formula
C mass formula
D reaction formula

8 **How many atoms** make up the compound shown below?

A 1
B 2
C 3
D 4

9 The name of the compound shown in the **chemical formula** below is _____.

A aluminum chlorine
B chlorine aluminide
C aluminum chloride
D chloroaluminum

10 What is the **correct** formula for **two** molecules of sodium bromide?

A NaBrx2
B 2NaBr
C NaBr2
D 2xNaBr