

**How many neutrons are in each of the following isotopes:**

Magnesium-24 has \_\_\_\_\_ neutrons

Radon-222 has \_\_\_\_\_ neutrons

Mercury 201 has \_\_\_\_\_ neutrons

Aluminum-27 has \_\_\_\_\_ neutrons

Osmium-190 has \_\_\_\_\_ neutrons

Platinum-195 has \_\_\_\_\_ neutrons

**Name the isotope with the following combination of protons and neutrons.**

58 protons and 82 neutrons = \_\_\_\_\_

53 protons and 74 neutrons = \_\_\_\_\_

20 protons and 20 neutrons = \_\_\_\_\_

82 protons and 125 neutrons = \_\_\_\_\_

**Working with ISOTOPES**

- Atomic Number = number of \_\_\_\_\_.
- Mass Number = number of \_\_\_\_\_ + \_\_\_\_\_ in an atom of an isotope
- number of neutrons = \_\_\_\_\_ Number – \_\_\_\_\_ Number.

Fill in the table below with the appropriate information for each **isotope**:

Isotope	Atomic Number	Mass Number	# of protons	# of neutrons
Germanium-73				
		16	8	
Gold-197				
		9	4	
Hydrogen-3				
Plutonium-244				
Potassium-39				
Bromine-80				

**Remember:** Atoms of the **SAME ELEMENT** with **DIFFERENT NUMBERS OF NEUTRONS** are called \_\_\_\_\_.