

**Drag and Drop**

salt	higher	low	7 or neutral	salt
corrosive	hydrogen ions	less or lower	dissolve	positive
solution	supersaturated	acids	water	hydroxide

35. A(n) \_\_\_\_\_ is a mixture that appears uniform throughout and whose particles cannot easily be separated.
36. When an ionic compound dissolves in water, the oxygen ends, **(which are negative)**, of water molecules are attracted to the \_\_\_\_\_ ions.
37. At room temperature, a \_\_\_\_\_ solution would contain a higher percent of solute than a saturated solution of the same substances.
38. The boiling point of a solution is \_\_\_\_\_ than that of the pure solvent.
39. Many bases can be recognized from their names, which often contain the word \_\_\_\_\_.
40. Substances that form hydrogen ions when released in water are called \_\_\_\_\_.
41. When the pH is high, **(a base)**, the concentration of hydrogen ions is \_\_\_\_\_.
42. In a reaction of an acid with a base, the pH changes to a value that is closer to \_\_\_\_\_.
43. Because it will dissolve so many solutes, \_\_\_\_\_ is sometimes called the universal solvent.
44. In a solution of salt water, the \_\_\_\_\_ is the solute.
45. Ionic and polar compounds \_\_\_\_\_ in water because water molecules are polar.
46. Strong acids produce more \_\_\_\_\_ than weak acids.
47. Foods such as oranges, tomatoes, and apples have a pH that is \_\_\_\_\_ than 7.
48. A(n) \_\_\_\_\_ is an ionic compound produced from the neutralization of an acid with a base.