

12. When a solute is added to a solvent, (ex. salt in water) the freezing point of the solution is

- higher than the freezing point of the pure solvent.
- lower than the freezing point of the pure solvent.
- the same as the freezing point of the pure solute.
- the same as the freezing point of the pure solvent.

13. Which is a characteristic property of acids?

- Acids turn blue litmus paper red.
- Acids turn red litmus paper blue.
- Acids taste bitter.
- Acids do not react with metals.

14. Strong acids can produce designs on metal printing plates because they

- turn red litmus paper blue.
- conduct electricity.
- are corrosive.
- react with limestone.

15. A substance that tastes bitter, feels slippery, and turns red litmus paper blue is a(n)

- acid.
- base.
- indicator.
- salt.

16. You are most likely to find a base in

- brick and metal cleaners.
- a car battery.
- fruit juice.
- household cleaners.

17. Any substance that forms hydrogen ions ( $H^+$ ) in water is a(n)

- acid.
- base.
- indicator.
- salt.

18. In water, bases form

- hydroxide ions.
- hydrogen ions.
- hydrogen gas.
- oxide ions.

19. Acids naturally present in food are safe to eat because they usually are

- concentrated.
- dilute.
- strong.
- weak.

20. The pH scale measures the

- strength of an acid.
- solubility of an acid.
- concentration of hydrogen ions.
- concentration of an acid.

21. You can use indicators to

- find the concentration of a solution.
- test for conductivity.
- find the pH of a solution.
- find out if a solution is saturated.

22. If a solution has a pH of 9, the solution is

- acidic.
- basic.
- neutral.
- saturated.

23. Normal rainfall is slightly acidic, which means its pH must be

- less than 2.
- between 5 and 7.
- between 2 and 4.
- between 7 and 9.

24. Neutralization is a reaction between

- an acid and a base.
- an acid and a metal.
- a base and a salt.
- salt and water.

25. A neutralization reaction produces

- a neutral solution.
- an acid and a base.
- water and a salt.
- table salt.