

NAME:

CLASS:

NUCLEAR NOTATION AND ISOTOPES

1) Complete the following table by typing the correct answers in the spaces provided. The first row is done as an example for you to follow.

You will need to use a Periodic Table to assist you in identifying the elements.

Atomic symbol	Name of Element	Mass Number	Atomic Number	No. of Protons	No. of Neutrons	No. of Electrons
${}^{48}_{22}\text{Ti}$	titanium	48	22	22	26	22
${}^{65}_{30}\text{Zn}$						
${}^{207}_{82}\text{Pb}$						
${}^{119}_{50}\text{Sn}$						
${}^{56}_{26}\text{Fe}$						
${}^{201}_{80}\text{Hg}$						

2) Complete the following sentences.

- Mass number is the sum of protons and _____ in the nucleus of an atom.
- Isotopes are _____ of the same element that have the same _____ number but the different _____ number, which means isotopes have the same numbers of _____ and electrons, but different numbers of _____.
- Isotopes have the same _____ properties but different _____ properties.