

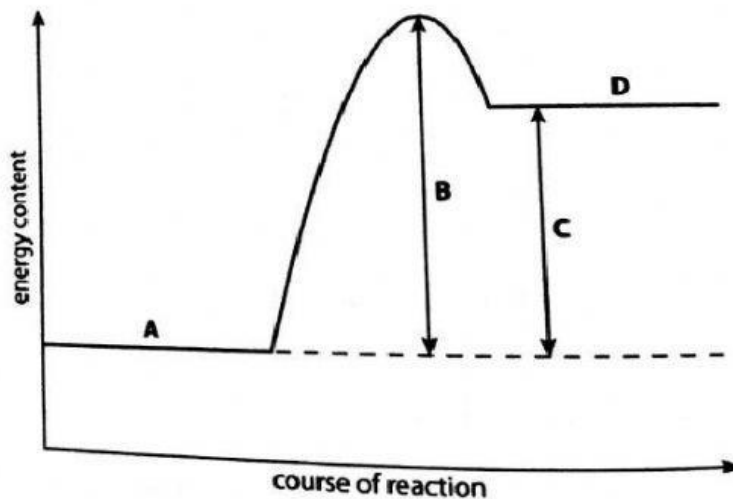
## BAHSE FORM FIVE CHEMISTRY ENERGETICS WORKSHEET

- 1) Complete the following table by choosing the correct word(s) from the brackets on the right.

For an ENDOTHERMIC REACTION		
a) Energy is _____ the surroundings.		<i>(lost to / gained from)</i>
b) $\Delta H$ is _____		<i>(positive / negative)</i>
c) The energy of products is _____ than the energy of the reactants		<i>(less / greater)</i>
d) The energy of bond formation is _____ than the energy of bond breaking.		<i>(less / greater)</i>
e) Temperature will _____		<i>(increase / decrease)</i>
For an EXOTHERMIC REACTION		
f) Energy is _____ the surroundings.		<i>(lost to / gained from)</i>
g) $\Delta H$ is _____		<i>(positive / negative)</i>
h) The energy of products is _____ than the energy of the reactants		<i>(less / greater)</i>
i) The energy of bond formation is _____ than the energy of bond breaking.		<i>(less / greater)</i>
j) Temperature will _____		<i>(increase / decrease)</i>

2) The diagram below shows the energy profile diagram for a reaction.

a) Identify A, B, C and D using the words in the Word Bank on the right.



WORD BANK
Activation energy
Enthalpy of reaction
Reactants
Products

A: \_\_\_\_\_

B: \_\_\_\_\_

C: \_\_\_\_\_

D: \_\_\_\_\_

b) In the diagram above, would  $\Delta H$  be **NEGATIVE** or **POSITIVE**? \_\_\_\_\_

c) In the diagram above, would the reaction be **ENDOTHERMIC** or **EXOTHERMIC**? \_\_\_\_\_