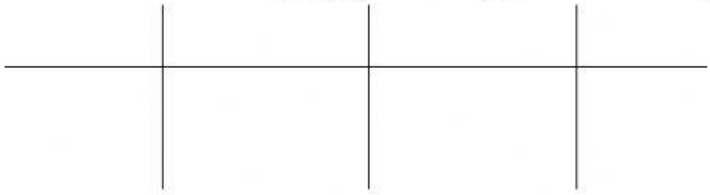
Name:_____

# STOICHIOMETRY - 1,2,3-Step

1. Calculate the number of **moles $CO_2$** produced by the reaction of **8.45 moles of C.**

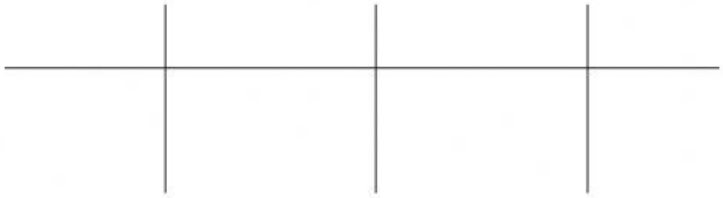
$2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$



=

2. Calculate the **grams of carbon** required to produce **5.67 g of iron.**

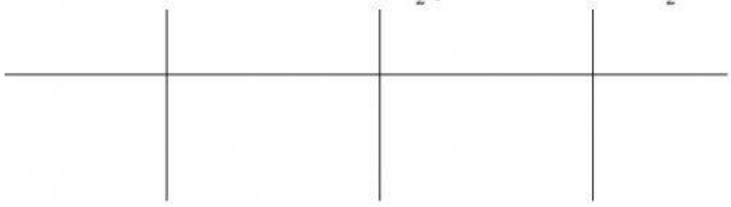
$2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$



=

3. Calculate the amount of **oxygen in grams** produced by the reaction of **125 L of water.**

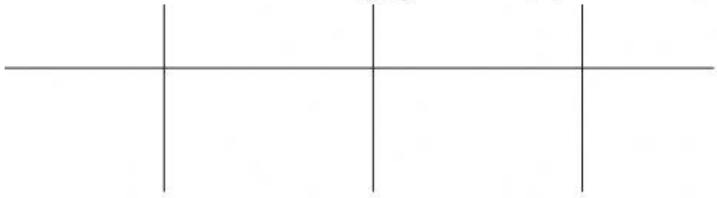
$2H_2O \rightarrow 2H_2 + O_2$



=

4. Calculate the number of **molecules of $CO_2$** produced by the reaction of **$2.45 \times 10^5$ representative particles of iron (III) oxide.**

$2Fe_2O_3 + 3C \rightarrow 4Fe + 3CO_2$



=

LIVEWORKSHEETS