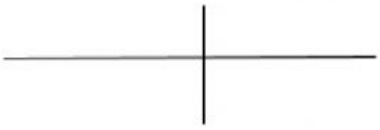
Name:
Stoichiometry

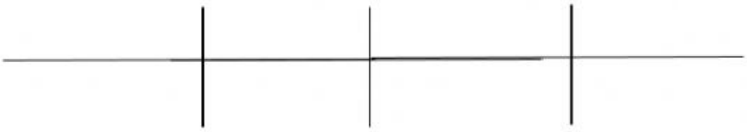
1. ___4___ **Fe** + ___3___ **$O_2$** → ___2___ **$Fe_2O_3$**

a) How many **moles** of Fe would be needed to react with 3.82 **moles** of $O_2$?



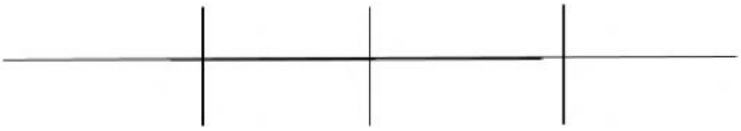
= mol Fe

b) What **mass** of $Fe_2O_3$ can be produced from 13.5 **grams** Fe?



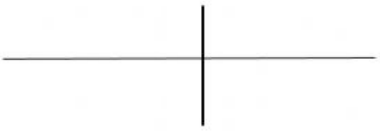
= g $Fe_2O_3$

c) How many **grams** of $O_2$ are needed to produce 34.7 **g** of $Fe_2O_3$?



= g $O_2$

d) What **moles** of $Fe_2O_3$ can be produced from 1.35 **mol** Fe?



= mol $Fe_2O_3$

2. ____$KClO_3$ ____KCl + ____$O_2$

When 62.0 g of Potassium chlorate decomposes, how many grams of KCl will be formed?

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LIVEWORKSHEETS