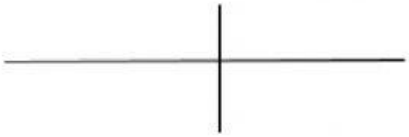
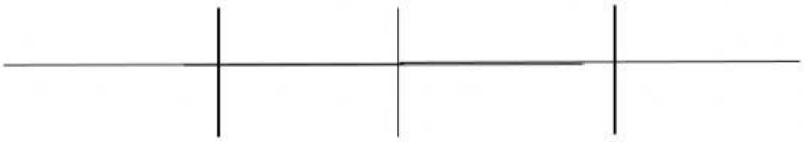
Name:
Stoichiometry

1. $\_\underline{4}\_\ \mathbf{Fe} + \_\underline{3}\_\ \mathbf{O_2} \rightarrow \_\underline{2}\_\ \mathbf{Fe_2O_3}$

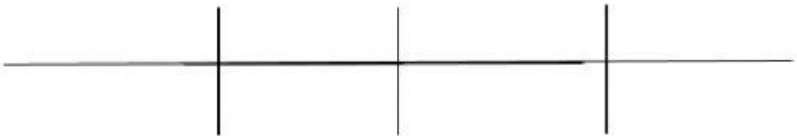
a) How many **moles** of Fe would be needed to react with 3.82 **moles** of $O_2$?

 = mol Fe

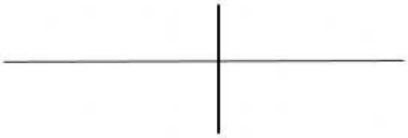
b) What **mass** of $Fe_2O_3$ can be produced from 13.5 **grams** Fe?

 = g $Fe_2O_3$

c) How many **grams** of $O_2$ are needed to produce 34.7 **g** of $Fe_2O_3$?

 = g $O_2$

d) What **moles** of $Fe_2O_3$ can be produced from 1.35 **mol** Fe?

 = mol $Fe_2O_3$

2. $\_\_\_\_KClO_3 \quad \_\_\_\_KCl + \_\_\_\_O_2$

When 62.0 g of Potassium chlorate decomposes, how many grams of KCl will be formed?

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LIVEWORKSHEETS