NAME: _____ BLOCK: _____
DATE: _____

CHEMISTRY: COUNTING ATOMS IN COMPOUNDS WORKSHEET #7.0.1

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: $5P_2O_3$ P = (5 × 2 =) 10 O = (5 × 3 =) 15

For example: $4Zn(NO_3)_2$ Zn = (4 × 1 =) 4 N = (4 × 1 × 2=) 8 O = (4 × 3 × 2 =) 24

1. $4K_2CO_3$ K = ____ C = ____ O = ____
2. $2Sr_3(PO_4)_2$ Sr = ____ P = ____ O = ____
3. $3N_4O_{10}$ N = ____ O = ____
4. $2(NH_4)_3N$ N = ____ H = ____
5. $8Cl_2O$ Cl = ____ O = ____
6. $Ca(C_2H_3O_2)_2$ Ca = ____ C = ____ H = ____ O = ____
7. $12NaBr$ Na = ____ Br = ____
8. $4Al(OH)_3$ Al = ____ O = ____ H = ____
9. $3NaHCO_3$ Na = ____ H = ____ C = ____ O = ____
10. $5Ga_2(Cr_2O_7)_3$ Ga = ____ Cr = ____ O = ____
11. $7C_2S_2$ C = ____ S = ____
12. $4Fe_2O_3$ Fe = ____ O = ____
13. $6Ba(MnO_4)_2$ Ba = ____ Mn = ____ O = ____
14. $3V_2O_5$ V = ____ O = ____
15. $2KNO_3$ K = ____ N = ____ O = ____
16. $9MgSO_4$ Mg = ____ S = ____ O = ____
17. $5Al_2(SiO_3)_2$ Al = ____ Si = ____ O = ____
18. $4Au(IO_3)_3$ Au = ____ I = ____ O = ____

(Continued)

LIVEWORKSHEETS

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: **$5P_2O_3$** **P = (5 × 2 =) 10** **O = (5 × 3 =) 15**

For example: **$4Zn(NO_3)_2$** **Zn = (4 × 1 =) 4** **N = (4 × 1 × 2=) 8** **O = (4 × 3 × 2 =) 24**

19. $8SnCl_4$ Sn = ____ Cl = ____
20. $6Cu_2SeO_4$ Cu = ____ Se = ____ O = ____
21. $3AsBr_3$ As = ____ Br = ____
22. $2H_2SO_4$ H = ____ S = ____ O = ____
23. SBr_2 S = ____ Br = ____
24. $4Ca(OH)_2$ Ca = ____ O = ____ H = ____
25. $5Mg_3(PO_4)_2$ Mg = ____ P = ____ O = ____
26. $12H_2O$ H = ____ O = ____
27. $5N_2O_4$ N = ____ O = ____
28. $3ClF$ Cl = ____ F = ____
29. $7P_2O_5$ P = ____ O = ____
30. $2KrCl_6$ Kr = ____ Cl = ____
31. $5Al(C_2H_3O_2)_2$ Al = ____ C = ____ H = ____ O = ____
32. $3(NH_4)_2Cr_2O_7$ N = ____ H = ____ Cr = ____ O = ____
33. $5Fe_3(PO_4)_2$ Fe = ____ P = ____ O = ____
34. $2NH_4NO_3$ N = ____ H = ____ O = ____
35. $5BaC_4H_4O_6$ Ba = ____ C = ____ H = ____ O = ____
36. $4Cu(HSO_3)_2$ Cu = ____ H = ____ S = ____ O = ____
37. $9Au(NO_2)_2$ Au = ____ N = ____ O = ____
38. $3K_2ZnO_2$ K = ____ Zn = ____ O = ____
39. $3Sr(MnO_4)_2$ Sr = ____ Mn = ____ O = ____
40. $4Al_2(CO_3)_3$ Al = ____ C = ____ O = ____

LIVEWORKSHEETS