

Persamaan Pengionan / Ionisation Equation

Ionisation of acid <i>Pengionan asid</i>			Ionisation of alkali <i>Pengionan alkali</i>		
HNO_3 Nitric acid <i>Asid nitrik</i>	$\longrightarrow$	 Hydrogen ion <i>Ion hidrogen</i>	+	 Nitrate ion <i>Ion nitrat</i>	
H_2SO_4 Sulphuric acid <i>Asid sulfurik</i>	$\longrightarrow$	 Hydrogen ion <i>Ion hidrogen</i>	+	 Sulphate ion <i>Ion sulfat</i>	
H_3PO_4 Phosphoric acid <i>Asid fosforik</i>	$\longrightarrow$	$3\text{H}^+(\text{aq/ak})$ Hydrogen ion <i>Ion hidrogen</i>	+	 Phosphate ion <i>Ion fosfat</i>	
CH_3COOH Ethanoic acid <i>Asid etanoik</i>	$\rightleftharpoons$	 Ethanoate ion <i>Ion etanoat</i>	+	 Hydrogen ion <i>Ion hidrogen</i>	
			Ammonia NH_3 $\xrightleftharpoons{\text{H}_2\text{O}}$ <i>Ammonia</i> Ammonium ion <i>Ion ammonia</i>		
			Sodium hydroxide $\text{NaOH} \longrightarrow$ <i>Natrium hidroksida</i> Sodium ion <i>Ion natrium</i>		
			Potassium hydroxide $\text{KOH} \longrightarrow$ <i>Kalium hidroksida</i> Potassium ion <i>Ion kalium</i>		
			Barium hydroxide $\text{Ba(OH)}_2 \longrightarrow$ <i>Barium hidroksida</i> Barium ion <i>Ion barium</i>		
			 Hydroxide ion <i>Ion hidroksida</i>		