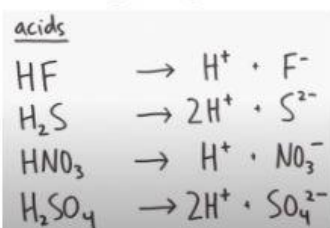


SC2. E. I can ask questions about chemical names to identify patterns in IUPAC nomenclature in order to predict chemical names for ionic, acidic, and inorganic covalent compounds.

Writing & Naming ACIDS Video Notes



1. An acid is a compound in which one or more \_\_\_\_\_ are bonded to a \_\_\_\_\_ ion.  
When naming acids just like in any ionic



2. Groups of elements that together have a charge are called \_\_\_\_\_ ions. Just like in any ionic formula the positive charge and the negative charge have to \_\_\_\_\_ out.
3. The name of an acid is based on the name of the \_\_\_\_\_ ion that is part of the acid.
4. There are two types of acids. One that have \_\_\_\_\_ and ones that don't have \_\_\_\_\_

<u>negative ion</u>	<u>acid</u>
-ide	hydro _____ ic acid
chloride	hydrochloric acid

**Naming Acids without Oxygen:** Write down what is shown in the video for this part.

HCl = \_\_\_\_\_ acid

HBr = \_\_\_\_\_ acid

SC2. E. I can ask questions about chemical names to identify patterns in IUPAC nomenclature in order to predict chemical names for ionic, acidic, and inorganic covalent compounds.

<u>negative ion</u>	<u>acid</u>
-ate	-ic acid
nitrate	nitric acid

**Naming Acids with Oxygen (Oxoacids):** Write down what is shown in the video for this part.

$\text{HNO}_3$  = \_\_\_\_\_ acid

$\text{H}_2\text{CO}_3$  = \_\_\_\_\_ acid

<u>negative ion</u>	<u>acid</u>
-ite	-ous acid
nitrite	nitrous acid

$\text{HNO}_2$  = \_\_\_\_\_ acid

$\text{HCrO}_2$  = \_\_\_\_\_ acid

<u>Polyatomic Ions</u>	
Carbonate	$\text{CO}_3^{2-}$
Chromite	$\text{CrO}_2^-$
Hypochlorite	$\text{ClO}^-$
Nitrate	$\text{NO}_3^-$
Nitrite	$\text{NO}_2^-$
Permanganate	$\text{MnO}_4^-$
Phosphate	$\text{PO}_4^{3-}$
Phosphite	$\text{PO}_3^{3-}$
Sulfate	$\text{SO}_4^{2-}$
Sulfite	$\text{SO}_3^{2-}$

<u>negative ion</u>	<u>acid</u>	<u>example</u>
-ide	hydro_____ic acid	chloride hydrochloric acid
-ate	-ic acid	nitrate nitric acid
-ite	-ous acid	nitrite nitrous acid

My r\_\_\_\_\_ has \_\_\_\_\_ l\_\_\_\_\_s.

I \_\_\_\_\_ something \_\_\_\_\_ ky.

Spr\_\_\_\_\_ is delici\_\_\_\_\_.

**Important Exception!**

<u>acids</u>	<u>negative ions</u>
$\text{H}_3\text{PO}_4$	phosphate ( $\text{PO}_4^{3-}$ )
$\text{H}_3\text{PO}_3$	phosphite ( $\text{PO}_3^{3-}$ )
$\text{H}_2\text{SO}_4$	sulfate ( $\text{SO}_4^{2-}$ )
$\text{H}_2\text{SO}_3$	sulfite ( $\text{SO}_3^{2-}$ )

$\text{HMnO}_4$  = \_\_\_\_\_ acid

$\text{HClO}$  = \_\_\_\_\_ acid

**So how do we name acids?**

Take a look at the \_\_\_\_\_ ion they have in them, we look at the \_\_\_\_\_ of that negative ion, and we use the \_\_\_\_\_ to figure out what we call the acid.

<u>Polyatomic Ions</u>	
Carbonate	$\text{CO}_3^{2-}$
Chromite	$\text{CrO}_2^-$
Hypochlorite	$\text{ClO}^-$
Nitrate	$\text{NO}_3^-$
Nitrite	$\text{NO}_2^-$
Permanganate	$\text{MnO}_4^-$
Phosphate	$\text{PO}_4^{3-}$
Phosphite	$\text{PO}_3^{3-}$
Sulfate	$\text{SO}_4^{2-}$
Sulfite	$\text{SO}_3^{2-}$

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