

INTRODUCTION TO REDOX CHEMISTRY

Complete the sentences below.

- Oxidation occurs when a substance _____ electrons and _____ in oxidation state.
- Reduction occurs when a substance _____ electrons and _____ in oxidation state.

Determine and then write the oxidation state of the underlined element in the space provided.

You must include the "+" or "-" sign in front of the oxidation number.

Example:



Oxidation State: **+6**

- $\underline{\text{S}}\text{O}_2$ Oxidation State: _____
- $\text{H}\underline{\text{N}}\text{O}_3$ Oxidation State: _____
- $\underline{\text{C}}\text{rO}_4^{2-}$ Oxidation State: _____
- $\text{Na}\underline{\text{C}}\text{lO}_3$ Oxidation State: _____
- $\underline{\text{M}}\text{nO}_2$ Oxidation State: _____
- $\text{V}\underline{\text{O}}_4^{3-}$ Oxidation State: _____
- $\underline{\text{M}}\text{nO}_4^{2-}$ Oxidation State: _____
- $\underline{\text{C}}\text{r}_2\text{O}_3$ Oxidation State: _____