

# Stability in Bonding

## Types of Bonds

Directions: Drag and Drop the words into the correct blank.

positive	element(s) <sub>1</sub>	compound(s) <sub>1</sub>	compound(s) <sub>2</sub>			
energy level(s)	charged	number(s)				
force(s) <sub>1</sub>	force(s) <sub>2</sub>	atom(s) <sub>1</sub>	atom(s) <sub>2</sub>	atom(s) <sub>3</sub>	ion(s)	negative
electron(s) <sub>1</sub>	electron(s) <sub>2</sub>	electron(s) <sub>3</sub>	electron(s) <sub>4</sub>			

1. A **chemical formula** tells what \_\_\_\_\_ make up a \_\_\_\_\_ and the exact \_\_\_\_\_ of atoms of each element in a unit of compound.
2. An atom is **chemically stable** when its outer \_\_\_\_\_ is completely filled with \_\_\_\_\_.
3. A **chemical bond** is a \_\_\_\_\_ that holds \_\_\_\_\_ together in a compound.
4. An \_\_\_\_\_ that has lost or gained \_\_\_\_\_ is called an **ion**.
5. An **ionic bond** is the \_\_\_\_\_ of attraction between the opposite charges of the \_\_\_\_\_ in an ionic \_\_\_\_\_.
6. The attraction that forms between \_\_\_\_\_ when they share \_\_\_\_\_ is known as a **covalent bond**.
7. A **polar molecule** has a slightly \_\_\_\_\_ + end and a slightly \_\_\_\_\_ - end.
8. A **nonpolar molecule** does not have oppositely \_\_\_\_\_ ends. Only atoms that are exactly alike can share their \_\_\_\_\_ equally.