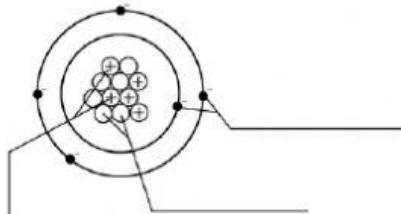


# Basic Atomic Structure

1. Label the parts of an atom in the diagram below



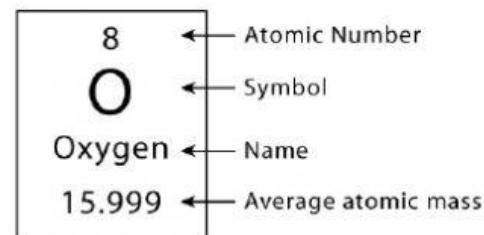
a) What type of charge does a proton have? \_\_\_\_\_

b) What type of charge does a neutron have? \_\_\_\_\_

c) What type of charge does an electron have? \_\_\_\_\_

d) Which two subatomic particles are located in the nucleus of an atom? \_\_\_\_\_

2. An element is represented by its chemical symbol along with a few numbers. Based on the example on the right, fill in the numbers of protons (P), neutrons (N), and electrons (E) for the following elements.



6
C
Carbon
12.011

#P \_\_\_\_\_  
#N \_\_\_\_\_  
#E \_\_\_\_\_

10
Ne
Neon
20.108

#P \_\_\_\_\_  
#N \_\_\_\_\_  
#E \_\_\_\_\_

19
K
Potassium
39.098

#P \_\_\_\_\_  
#N \_\_\_\_\_  
#E \_\_\_\_\_

3. Complete the table using your knowledge of the periodic table.

Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of Neutrons
Na			11		12
K		39		19	
			38		50
F				9	10
	20	40		20	
	50			50	69
I	53	127			