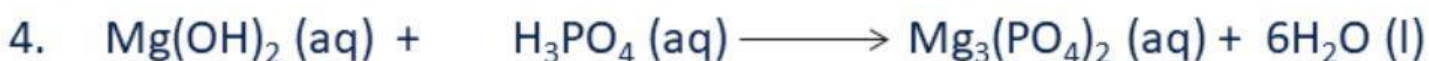


Nama: \_\_\_\_\_

Balance the following chemical equations:

1



Calculate the oxidation number for the underlined atom:

	FORMULA	OXIDATION NUMBER OF UNDERLINED ATOM
1.	<u>Cl</u>	
2.	<u>K</u>	
3.	<u>F</u>	
4.	<u>Al</u> <sup>3+</sup>	
5.	<u>O</u> <sub>2</sub>	

	FORMULA	OXIDATION NUMBER OF UNDERLINED ATOM
6.	<u>K</u> <sup>+</sup>	
7.	<u>N</u> O <sub>3</sub> <sup>-</sup>	
8.	<u>S</u> O <sub>4</sub> <sup>2-</sup>	
9.	<u>Mn</u> O <sub>2</sub>	
10.	<u>Fe</u> O	

	FORMULA	OXIDATION NUMBER OF UNDERLINED ATOM
11.	H <sub>2</sub> <u>S</u>	
12.	<u>P</u> O <sub>4</sub> <sup>3-</sup>	
13.	<u>Si</u> O <sub>2</sub>	
14.	H <sub>2</sub> <u>O</u> <sub>2</sub>	
15.	<u>N</u> H <sub>4</sub> Cl	

In a free element, as an atom or a molecule the oxidation number is zero

For monoatomic ion, the oxidation number is equal to the charge on the ion

Fluorine and other halogens have oxidation number of -1 except when combine with oxygen

Hydrogen is +1 except in metal hydrides (-1)

Oxygen is -2 except in peroxide (-1) and when combine with fluorine (+2)

In neutral molecule, the sum of the oxidation number of all atoms that made up the molecule is equal to zero

For polyatomic ions, the total oxidation number of all atoms that made up the polyatomic ion must be equal to the net charge of the ion