

Learning Target: *I can use the International Union of Pure and Applied Chemistry (IUPAC) nomenclature for translating between simple binary chemical names and chemical formulas.*

Naming Binary Covalent & Ionic Chemical Formulas & Compounds Interactive Activity

THE FIRST THING YOU DO WHEN NAMING FORMULAS AND COMPOUNDS IS DETERMINE IF IT IS AN _____. IF IT IS A COVALENT BOND THEN IT CONTAINS _____. IF IT IS AN IONIC BOND THEN IT CONTAINS _____. YOU NEED TO KNOW THE _____ FOR IONIC COMPOUNDS AND _____ THE CHARGES FOR IONIC CHEMICAL FORMULAS. YOU NEVER PUT _____ FOR IONIC BONDS. YOU NEED TO KNOW THE _____ FOR COVALENT COMPOUNDS. YOU NEVER PUT _____ IN FRONT OF THE FIRST WORD OF A COVALENT COMPOUND. YOU ADD _____ AT THE END OF THE SECOND WORD FOR IONIC AND COVALENT COMPOUNDS.

1. What is an ionic bond made up of? What is a covalent bond made up of?
2. What do you need to name ionic bonds? What do you need to name covalent bonds?
3. What do you put at the end of both ionic and covalent bonds?
4. What do you do with the charges of an ionic bond? What do ionic bonds NOT need?
5. How do you name the 1st word of an ionic compound? 2nd word of an ionic compound?
6. How do you name the 1st word of a covalent compound? 2nd word of a covalent compound?
7. What number do you never add as a subscript?

Naming Ionic Compounds Practice:

MgF ₂	ZnCl ₂	K ₂ O	Al ₂ O ₃	Na ₂ S
Aluminum Sulfide	Iron (III) Bromide	Lithium Oxide		
Beryllium Chloride	Potassium Oxide			

Naming Covalent Compounds Practice:

IF ₅	SbBr ₃	N ₂ O ₃	B ₆ Si	ClO ₂
Hydrogen Iodide	Phosphorus Triiodide	Sulfur Tetrachloride		
Nitrogen Dioxide	Disilicon Hexabromide			

Scan the QR Code
to take the Quiz!

