Roatan Bilingual School

10th Grade Chemistry Evaluation

SEMESTER RECUPERATION

Name: _____ Date: _____

Ms. April Brown 100%

Read each question or exercise carefully. RESPOND IN YOUR OWN WORDS. to show your work and leave all evidence possible.

TRUE OR FALSE

Read each statement carefully. Write a T on the line next to it if it is true. Write an F if it is false and EXPLAIN WHY.

1. ____ The formula for Carbon Monoxide is CO_2.
2. ____ KBr is Potassium Baromdiide.
3. _____ NaF is a covalent compound.
4. _____ ZnI is an ionic compound.
5. ____ Pressure is a force or characteristic that can be measured in gasses.

   The unit used to measure pressure is calories.
6. _____ H2O is a molecule.
7. _____ Ne is a noble gas.
8. _____ H2O and N2 are examples of matter that are in plasma state.
9. _____ Temperature is measured ONLY in Celsius.
10. _____ Mass is usually measured in L.
11. _____ CaCO represents Calcium Nitrate
12. _____ Calories are used to measure temperature.

LIVEWORKSHEETS

MULTIPLE CHOICE

Circle the best answer. If your choice is ¨none of the above¨ explain why.

1. **This is a polyatomic compound**
   a. Acetate
   b. $NaNO_3$
   c. K_3PO_4
   d. All of the above
2. **Name KOH**
   a. Cerium Oxide
   b. Kallium Oxihydride
   c. Potassium Hydroxide
   d. None of the above _____
3. **What does the subscript in $Al_2(SO_4)_3$ represent?**
   a. There is no subscript
   b. This is a formula for a different compound
   c. This is a covalent compound
   d. None of the above _____
4. **Equation for a chemical reaction in which the number of atoms for each element in the reaction and the total charge is the same for both the reactants and the products**
   a. Potassium, Copper, Carbon
   b. Calcium, Magnesium, Silver
   c. synthesis reaction
   d. None of the above _____
5. **Avogadro´s number**
   a. 1 x 10 E23
   b. 6.02 x 10 E23
   c. 62.0 x 10 E23
   d. None of the above _____
6. **Unit used to measure the amount or quantity of substances in a chemical reaction.**
   a. Liters
   b. Moles
   c. Atoms
   d. None of the above _____
7. **(2.016 g/mol) is the molar mass of:**
   a. N2
   b. H2

LIVEWORKSHEETS

c. Cl2

d. None of the above _____

8. **A state of matter in which particles are NOT tightly packed together.**

a. liquid

b. plasma

c. solid

D. None of the above _____

9. **In the equation CO2 + H2O + ATP —> O2 + C6H12O2, which compound is part of the product?**

a. H2O

b. CO2

c. ATP

D. None of the above _____

9. **In the equation CO2 + H2O + ATP —> O2 + C6H12O6, which compound is part of the reactants?**

a. H2O

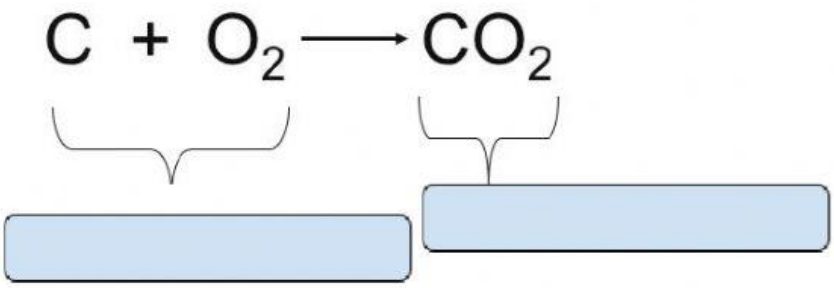
b. C6H12O6

c. O2

D. None of the above _____

PRACTICAL EXERCISES

1. Name the parts of an equation



2. Balance the equations

$H_2 + O_2 \longrightarrow H_2O$

$HCL + 2Na \longrightarrow NaCl + H_2$

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3. Write a synthesis reaction: N, H, $2NH_3$ (balance the equation)

4. Write a decomposition reaction: Zn, S, ZnS

5. Name the compounds

   $Mg_3(PO_4)_2$

   NH3

   $C_6H_{12}O_6$

   $Al_2(SO_4)_3$

6. Write the formula

   Glucose=

   Lead Nitrate

   Calcium Hydroxide

7. Tell whether single displacement or double displacement

   $H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2H_2O$

   $Zn + 2HCl \rightarrow H_2 + ZnCl_2$

8. Point out the subscripts.

$$3Fe_{(s)} + 4H_2O_{(g)} \longrightarrow Fe_3O_{4(s)} + 4H_{2(g)}$$

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Reactants

Iron

Hydrogen

Oxygen

Products

Iron

Oxygen

Hydrogen

9. Point out the coefficients



Reactants

Iron

Hydrogen

Oxygen

Products

Iron

Oxygen

Hydrogen

1. Describe the 4 states of matter discussed in class. 1pt

2. Name the following compounds. 4pts

Ca(OH)

HCL

NaCl

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3. Convert the temperatures. Use the conversion factors. 2pts
110F to C

50C to F

4. Convert the amounts. Use the conversion factors. 3pts

1 L = 1000mL 1 kg = 1000g 1hr = 60min

360min to hr

600 mL to L

8900g to Kg

CELSIUS TO FAHRENHEIT

$$T_F = \left(\frac{9}{5}T_C\right) + 32$$

FAHRENHEIT TO CELSIUS

$$T_C = \frac{5}{9}\left(T_F - 32\right)$$

5. Balance the equations. 2pts

_____ CH_2O + _____ H_2 → _____ CH_3OH

_____ P_4 + _____ H_2 → _____ PH_3

6. Find the molar mass. 4pts

Ethane C_2H_6

KCl

7. Covert from moles to grams. 2pts

7.5 moles of CO_2

10 moles of H_2

8. Convert from grams to moles. 2pts

454 g of NH_3

70g of NO_2

9. Identify the formula units. 2pts

0.533 mol NH_4

1. 65 g of H_2

LIVEWORKSHEETS