31. Count the atoms in the formulas: **$C_7H_5(NO_2)_3$**, C - _____ H- _____ N - _____ O - _____

$Ca(H_2PO_4)_2$ Ca - _____ H - _____ P - _____ O - _____

a. **$C_7H_5(NO_2)_3$ and $Ca(H_2PO_4)_2$** are examples of a(n) (**element, compound, mixture, abbreviation.**)

32. Label the box from the periodic table:

a. 12 _____

b. Mg_____

c. Magnesium_____

d. 24.305_____

12
Mg
Magnesium
24.305

Average Atomic Mass

Element Symbol

Element Name

Atomic Number

Drag and Drop Properties

Not Shiny
Malleable
Not Ductile
Brittle
Electricity
Electricity
Liquid
Shiny
Not Malleable
Ductile
Hardness
Heat
Heat
Solid
Gas

33. The properties of metals are:

a. _____

b. _____

c. _____

d. _____

e. Good conductors of _____and _____

f. Most of them are in the _____ state at room temperature.

34. The properties of non-metals are:

a. _____

b. _____

c. _____

d. _____

e. Poor conductors of _____and _____

f. Most of them are in the _____ state at room temperature.

35. All of the Laws of Conservation state that nothing is _____ or destroyed, _____ or _____.

36. Which of these cannot be broken down in to another substance by ordinary chemical or physical means: (**a molecule, a mixture, an element, an acid like HCl**).

37. Choose Compound, mixture, or element

a. Salt water _____

b. Oxygen (O) _____

c. Air _____

d. Table Salt (NaCl) _____

LIVEWORKSHEETS