

**Q. 1 Match the correct pairs : (Use the pencil and draw lines)****Element [protons]****Electronic Configuration****1. Magnesium[p=12]****a) 2, 8, 3****2. Sulphur[p=16]****b) 2, 8****3. Neon[p=10]****c) 2, 8, 2****4. Calcium[p=20]****d) 2, 8, 6****5. Aluminium [p=13]****e) 2, 8, 8, 2****Q. 2 Choose the word options and drag them into the correct places:** **$2n^2$** **Mass No.****Valency****Atomic No.**

1- total number of protons and neutrons present in a neutral atom.

2-It is formula to determine the maximum number of electrons in each shell of an atom.

3-The number of electrons or protons that are present in a neutral atom.

4-the number of electrons lost or gained by an atom to achieve a stable electronic configuration.

Q. 3 Drag the electron/s from the first figure and drop it into the second figure to show the electron transfer and then answer the questions given below:

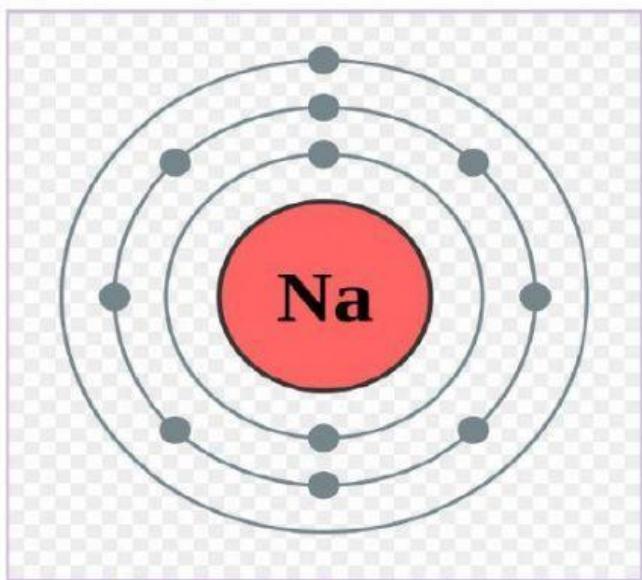


Fig 1 : SODIUM

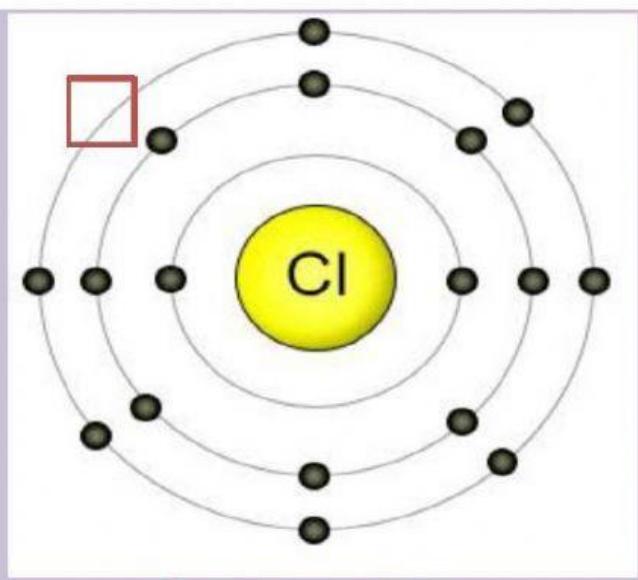


Fig 2 : CHLORINE

1. What kind of bond is achieved by this method?

2. Which of these two is a metal?

3. Which is the nearest inert gas to Chlorine?

4. What is the valency of sodium?

5. What is the valency of Chlorine?

Q. 4 Match the element symbol with its valency : (Draw lines)

