



**Q. 1 Match the correct pairs : ( Use the pencil and draw lines )**

**Element [protons]**

**Electronic Configuration**

1. Magnesium[p=12]

a) 2, 8, 3

2. Sulphur[p=16]

b) 2, 8

3. Neon[p=10]

c) 2, 8, 2

4. Calcium[p=20]

d) 2, 8, 6

5. Aluminium [p=13]

e) 2, 8, 8, 2

**Q. 2 Choose the word options and drag them into the correct places:**

$2n^2$

Mass No.

Valency

Atomic No.

1- total number of protons and neutrons present in a neutral atom.

2-It is formula to determine the maximum number of electrons in each shell of an atom.

3-The number of electrons or protons that are present in a neutral atom.

4-the number of electrons lost or gained by an atom to achieve a stable electronic configuration.

**Q. 3 Drag the electron/s from the first figure and drop it into the second figure to show the electron transfer and then answer the questions given below:**

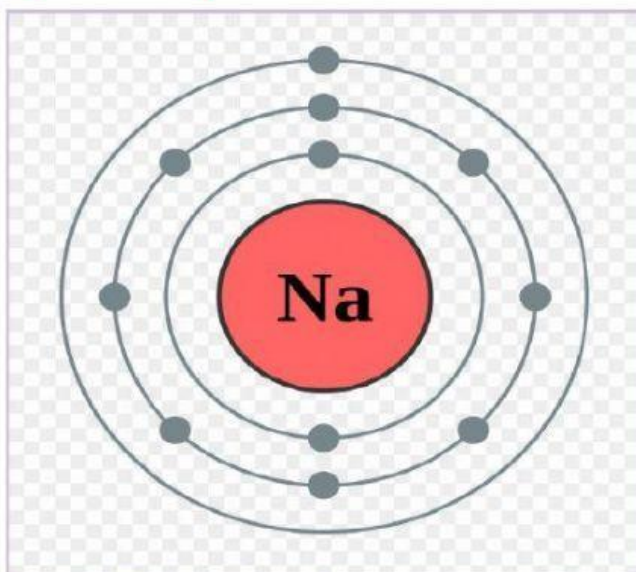


Fig 1 : SODIUM

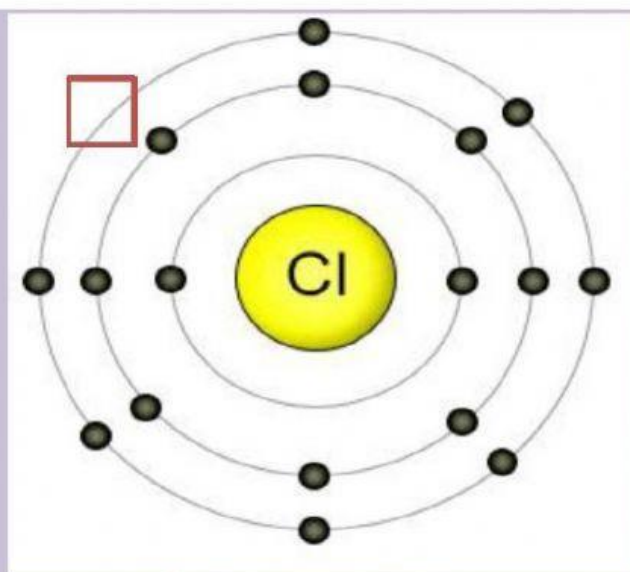


Fig 2 : CHLORINE

1. What kind of bond is achieved by this method?
2. Which of these two is a metal?
3. Which is the nearest inert gas to Chlorine?
4. What is the valency of sodium?
5. What is the valency of Chlorine?

**Q. 4 Match the element symbol with its valency : (Draw lines )**