

## Ionic vs Molecular Compounds – A Review

Ionic compounds form by transferring electrons between metals and nonmetals. Molecular or covalent compounds form by sharing electrons between nonmetal elements.

Identify the following as Ionic (I) or Molecular (M)

Compound	Ionic (I) or Molecular (M)
$\text{SF}_6$	
dinitrogen monoxide	
$\text{KBr}$	
$\text{CaI}_2$	
$\text{P}_4\text{O}_6$	
magnesium sulfate	
sodium hydroxide	
carbon tetrachloride	
$\text{Na}_2\text{SO}_3$	
$\text{Cl}_2\text{O}_7$	
$\text{NaHCO}_3$	
$\text{N}_2\text{O}_3$	
tetraphosphorus hexoxide	
mercury(I) oxide	
boron trichloride	
iron(III) sulfide	