

SPM 2012 Soalan 6

4 Antara zarah yang berikut, yang manakah bersamaan dengan 1 mol?

Which of the following particles equal to 1 mole?

- A Bilangan atom dalam 1 g gas hidrogen
The number of atom in 1 g of hydrogen gas
- B Bilangan molekul dalam 1 g gas hidrogen
The number of molecule in 1 g of hydrogen gas
- C 6.02×10^{23} atom hidrogen dalam gas hidrogen
 6.02×10^{23} of hydrogen atoms in hydrogen gas.
- D 6.02×10^{23} molekul hidrogen dalam gas hidrogen
 6.02×10^{23} of hydrogen molecule in hydrogen gas.

SPM 2007 Soalan 36

5 Antara gas berikut, yang manakah mengandungi 0.4 mol atom pada suhu dan tekanan bilik?

[1 mol gas menempati isi padu sebanyak 24 dm³ pada suhu dan tekanan bilik]

Which of the following gases contains 0.4 mol of atoms at room temperature and pressure?

[1 mol of gas occupies the volume of 24 dm³ at room temperature and pressure]

- A 4.8 dm³ He
- B 4.8 dm³ H₂
- C 4.8 dm³ SO₃
- D 4.8 dm³ CO₂

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6 Apabila kuprum(II) karbonat, CuCO₃ dipanaskan, gas yang terbebas menukarkan air kapur menjadi keruh.

Berapakah isi padu gas yang terbebas apabila 0.62 g kuprum(II) karbonat dipanaskan pada keadaan bilik?

[Jisim atom relatif ; C = 12, O = 16, Cu = 64 ; Isipadu molar gas = 24 dm³ mol⁻¹ pada keadaan bilik]

When copper(II) carbonate, CuCO₃ is heated, the gas released turns the lime water chalky.

What is the volume of gas released when 0.62 g of copper(II) carbonate is heated at room condition?

[Relative atomic mass; C = 12, O = 16, Cu = 64; Molar volume of gas = 24 dm³ mol⁻¹ at room conditions]

- A 5 cm³
- B 120 cm³
- C 240 cm³
- D 360 cm³

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7. Formula bagi ion sulfat adalah SO₄²⁻ dan ion nitrat adalah NO₃⁻.

Jika formula garam sulfat bagi M ialah MSO₄, apakah formula garam nitrat bagi M?

The formula for a sulphate ion is SO₄²⁻ and for a nitrate is NO₃⁻.

If the formula of the sulphate salt of M is MSO₄, what is the formula of the nitrate salt of M?

- A MNO₃
- B M₂NO₃
- C M(NO₃)₂
- D M(NO₃)₃

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8 Sebatian manakah yang mempunyai formula yang betul?

Which compound has the correct formula?

	Sebatian Compound	Formula Formula
A	Barium Nitrat <i>Barium Nitrate</i>	Ba(NO ₃) ₂
B	Plumbum(II) oksida <i>Lead(II) oxide</i>	PbO ₂
C	Kuprum(II) oksida <i>Copper(II) oxide</i>	Cu ₂ O
D	Argentum karbonat <i>Silver carbonate</i>	AgCO ₃

SPM 2009 Soalan 46

9 Persamaan kimia berikut menunjukkan pembakaran lengkap bagi gas etana

The following chemical equation shows the complete combustion of ethane gas.



Berapakah isipadu etana yang diperlukan untuk menghasilkan 2.2 g karbon dioksida pada suhu dan tekanan piawai?

[Jisim molar CO₂ = 44 gmol⁻¹. Isipadu molar gas pada suhu dan tekanan piawai = 22.4 dm³mol⁻¹]

What is the volume of ethane needed to produce 2.2 g of carbon dioxide at standard temperature and pressure?

[*Molar mass of CO₂ = 44 gmol⁻¹. Molar volume of gas at standard temperature and pressure = 22.4 dm³mol⁻¹*]

- | | | | |
|---|----------------------|---|----------------------|
| A | 0.05 dm ³ | B | 0.10 dm ³ |
| C | 0.56 dm ³ | D | 1.12 dm ³ |

SPM 2010 Soalan 4

10 Persamaan berikut mewakili satu tindak balas .

The following equation represents a reaction.



Apakah bahan – bahan tindak balas dalam persamaan ini?

What are the reactants in this equation?

- A Kuprum(II) nitrat dan air
Copper(II) nitrate and water
- B Kuprum(II) nitrat dan asid nitrik
Copper(II) nitrate and nitric acid
- C Kuprum(II) hidroksida dan asid nitrik
Copper(II) hydroxide and nitric acid
- D Kuprum(II) hidroksida dan kuprum(II) nitrat
Copper(II) hydroxide and copper(II) nitrate.

