SALTS WORKSHEET – FORM FOUR CHEMISTRY

NAME:

1. Complete the table to give information about the solubility of EACH compound by writing SOLUBLE or INSOLUBLE.

COMPOUND	**SOLUBILITY**
sodium carbonate	
copper (II) nitrate	
lead (II) sulphate	
calcium carbonate	
iron (III) hydroxide	
silver chloride	
magnesium sulphate	
potassium hydroxide	
ammonium chloride	
barium sulphate	
sodium sulphate	
lead (II) iodide	

LIVEWORKSHEETS

2. Use the word list below to answer the questions that follow

TITRATION	**NEUTRALISATION**	**PRECIPITATION**
ACID	**ANTACID**	**SULPHURIC ACID**
FILTRATION	**EVAPORATION**	**CRYSTALLISATION**
INSOLUBLE	**NORMAL**	**SOLUBLE**

(i) Alex wants to prepare a sample of copper (II) sulphate in the laboratory, starting with copper (II) carbonate. Name the other reactant he would use.

(ii) Alex adds the copper (II) carbonate to the reactant in excess to ensure complete reaction. What method would he use to obtain the copper (II) sulphate solution from the excess copper (II) carbonate?

(iii) Simon wants to prepare a sample of sodium chloride. What technique would he use?

(iv) Michael want to prepare lead (II) carbonate in the lab. What technique would he use?

This because lead (II) carbonate is:

(v) A/an _____ salt is formed when some of the H+ ions in an acid are replaced by metal ions.

(vi) A/an _____ salt is formed when all of the H+ ions in an acid are replaced by metal ions.

(vii) A/an _____ is a substance used to neutralise stomach acid.

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