

Nama :

#### UJI KEMAMPUAN DIRI BAB IKATAN KIMIA

1. Ikatan ion disebabkan oleh adanya...
  - a. Pemakaian bersama sepasang elektron
  - b. Gaya tarik menarik antara kation dengan anion
  - c. Ikatan antara unsur nonlogam
  - d. Pemakaian elektron berasal dari satu pihak
  - e. Ikatan antara unsur-unsur yang cenderung menarik elektron
2. Atom  ${}_{20}\text{A}$  dan  ${}_{17}\text{B}$  akan membentuk senyawa yang ....
  - a. Berikatan ion dengan rumus senyawa  $\text{AB}_2$
  - b. Berikatan ion dengan rumus senyawa  $\text{A}_2\text{B}$
  - c. Berikatan ion dengan rumus kimia  $\text{AB}$
  - d. Berikatan kovalen dengan rumus kimia  $\text{AB}_2$
  - e. Berikatan kovalen dengan rumus kimia  $\text{A}_2\text{B}$
3. Diketahui beberapa unsur sebagai berikut :  
 ${}_{3}\text{A}$ ,  ${}_{4}\text{B}$ ,  ${}_{9}\text{C}$  dan  ${}_{10}\text{D}$   
Ikatan ion dapat terbentuk antara...
  - a. A dengan B
  - b. B dengan C
  - c. C dengan D
  - d. D dengan A
  - e. B dengan D
4. Rumus senyawa yang terbentuk dari  ${}_{13}\text{Al}$  berikatan dengan  ${}_{17}\text{Cl}$  adalah ...
  - a.  $\text{AlCl}$
  - b.  $\text{AlCl}_2$
  - c.  $\text{AlCl}_3$
  - d.  $\text{Al}_2\text{Cl}$
  - e.  $\text{Al}_3\text{Cl}$
5. Pernyataan berikut yang benar tentang ikatan kovalen adalah ...
  - a. Adanya pemakaian pasangan bersama pasangan elektron yang berasal dari kedua atom yang berikatan
  - b. Terjadi antara atom-atom yang memiliki energi ionisasi rendah dengan atom yang memiliki afinitas elektron yang tinggi
  - c. Terjadi gaya tarik menarik antara muatan yang berbeda
  - d. Terbentuk dari unsur logam dan non logam
  - e. Terjadi akibat serah terima elektron dari atom yang satu ke atom yang lain

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6. Di antara senyawa di bawah ini, yang berikatan kovalen adalah...

- a. NaCl
- b.  $\text{NH}_3$
- c. CaO
- d.  $\text{MgF}_2$
- e. KCl

7. Di antara molekul-molekul di bawah ini yang mempunyai ikatan kovalen rangkap tiga adalah...

- a.  $\text{N}_2$
- d.  $\text{H}_2\text{O}$
- b.  $\text{H}_2$
- e.  $\text{NH}_3$
- c.  $\text{O}_2$

**Partial Periodic Table of the Elements**

1 H Hydrogen 1.00794	2 He Helium 4.0026
3 Li Lithium 6.941	4 Be Beryllium 9.0122
11 Na Sodium 22.9898	12 Mg Magnesium 24.3050
19 K Potassium 39.0983	20 Ca Calcium 40.078

13 B Boron 10.811	14 C Carbon 12.0107	15 N Nitrogen 14.0067	16 O Oxygen 15.9994	17 F Fluorine 18.9984	18 Ne Neon 20.1797
13 Al Aluminum 26.98154	14 Si Silicon 28.0855	15 P Phosphorus 30.9738	16 S Sulfur 32.065	17 Cl Chlorine 35.4527	18 Ar Argon 39.948

The diagram illustrates the periodic table with elements 1 through 20 highlighted in a grid. The elements are arranged in four rows, with the first two rows containing elements 1-10 and the next two rows containing elements 11-20. The elements are labeled with their atomic number, symbol, name, and atomic weight. The grid is divided into two main sections: the first section (left) contains elements 1-10, and the second section (right) contains elements 11-20. The elements are arranged in a grid that is 4 rows high and 2 columns wide. The first row contains elements 1 (H) and 2 (He). The second row contains elements 3 (Li) and 4 (Be). The third row contains elements 11 (Na) and 12 (Mg). The fourth row contains elements 19 (K) and 20 (Ca). The elements are arranged in a grid that is 4 rows high and 2 columns wide. The first row contains elements 1 (H) and 2 (He). The second row contains elements 3 (Li) and 4 (Be). The third row contains elements 11 (Na) and 12 (Mg). 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8. Pasangan senyawa yang keduanya memiliki ikatan kovalen adalah...

- a. NaCl dan  $\text{H}_2\text{S}$
- b. HCl dan  $\text{NH}_3$
- c. BaO dan  $\text{AlCl}_3$
- d. NaI dan CaO
- e.  $\text{CaCl}_2$  dan BaS

9. Data konfigurasi elektron dari beberapa unsur sebagai berikut.

P = 2 8 1

Q = 2 8 5

R = 2 8 8

S = 2 8 7

Pasangan unsur yang dapat membentuk ikatan kovalen adalah...

- a. P dan Q
- b. P dan S
- c. Q dan S
- d. Q dan R
- e. P dan R

10. Pasangan molekul di bawah ini yang berikatan kovalen polar adalah ...

- a.  $\text{H}_2\text{O}$  dan  $\text{O}_2$
- b.  $\text{Cl}_2$  dan HCl
- c.  $\text{CO}_2$  dan  $\text{O}_2$
- d.  $\text{Cl}_2$  dan  $\text{N}_2$
- e. HCl dan  $\text{H}_2\text{O}$