

PRACTICE CHAPTER 5 & 6

Identify True (✓) or False Statements (X) .

1. The following sentences are about the assumptions made on **Kinetic Molecular Theory of Ideal**

Gas :

(i) gas particles are extremely small compared to the large spaces in the container.	TRUE	FALSE
(ii) gas particles attract and repulse one another .	TRUE	FALSE
(iii) collision among gas particles are elastic and the motions are random.	TRUE	FALSE
(iv) When temperature increases, kinetic energy of particles decrease.	TRUE	FALSE
2. Real gases behave ideally when temperature is low and pressure is high.	TRUE	FALSE
3. Liquids with strong intermolecular forces have low vapour pressure and low boiling point.	TRUE	FALSE
4. Water beads on the surface of leaf after rainfall because of the property of liquid which is viscosity.	TRUE	FALSE
5. Viscosity increases as temperature increases.	TRUE	FALSE
6. Viscosity decreases when liquid has strong intermolecular forces.	TRUE	FALSE
7. Liquid boils when vapour pressure of the liquid equal the external atmospheric pressure.	TRUE	FALSE
8. Metallic solids have valence electrons that are mobile. (can move or delocalize)	TRUE	FALSE
9. The process in which a gas transform into liquid is called vaporization.	TRUE	FALSE
10. The following sentences describe the characteristics of a dynamic equilibrium:		
(i) The concentration of products and reactants are the same.	TRUE	FALSE
(ii) Rate of forward reaction is not equal rate of reverse reaction.	TRUE	FALSE
(iii) There is no net change in the system.	TRUE	FALSE
(iv) The concentration of products and reactants have stopped changing.	TRUE	FALSE
11. A reaction is at equilibrium when the value of Q_c is equal to K_c .	TRUE	FALSE
12. If Q_c is less than K_c , that means there are more products in the mixture.	TRUE	FALSE
13. When Q_c is not equal to K_c , the reaction will proceed until $Q_c = K_c$.	TRUE	FALSE
14. Applying Le Chatelier's principle, a change in pressure has no effect on the equilibrium position if there are same number of moles of reacting gases on both sides of equation.	TRUE	FALSE