Question 5 b (iii)

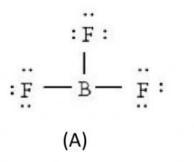
Predict the molecular geometry, bond angle, polarity and type of IMF of the BF₃

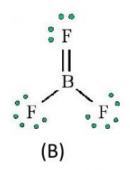
B in Group 13

Calculate the formal charge of each atom

F in Group 17

and determine the correct Lewis structure of BF₃_____





· Electron pair arrangement at central atom B:

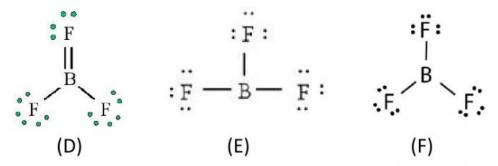
____ bonding pairs electrons.

Basic shape is ______

• VSEPR: The **repulsion** between bonding pairs electrons is

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 State the shape of molecule _____ and choose the correct molecular geometry of BF₃ _____



- Every F-B-F bond angle is _______°
- _____is more electronegative than _____



- Dipole moment can _____ each other.
- Net dipole moment (μ _____ 0)
- Therefore, it is a _____ molecule.
- Intermolecular forces in BF₃: _____

Note:

Hydrogen bond = HB

Dipole-dipole force = DDF (for polar compound)

London Dispersion forces = LDF (for non-polar compound)

