Average atomic mass

$$Average\ atomic\ mass = \frac{\%\ of\ abundance}{100}\ X\ atomic\ mass$$

Isotope	% abundance	mass number
$^{6}_{12}C$	98.89%	12
$^{6}_{13}C$	1.11%	13

Average Atomic mass = $\frac{98.89}{100} \times 12 +$

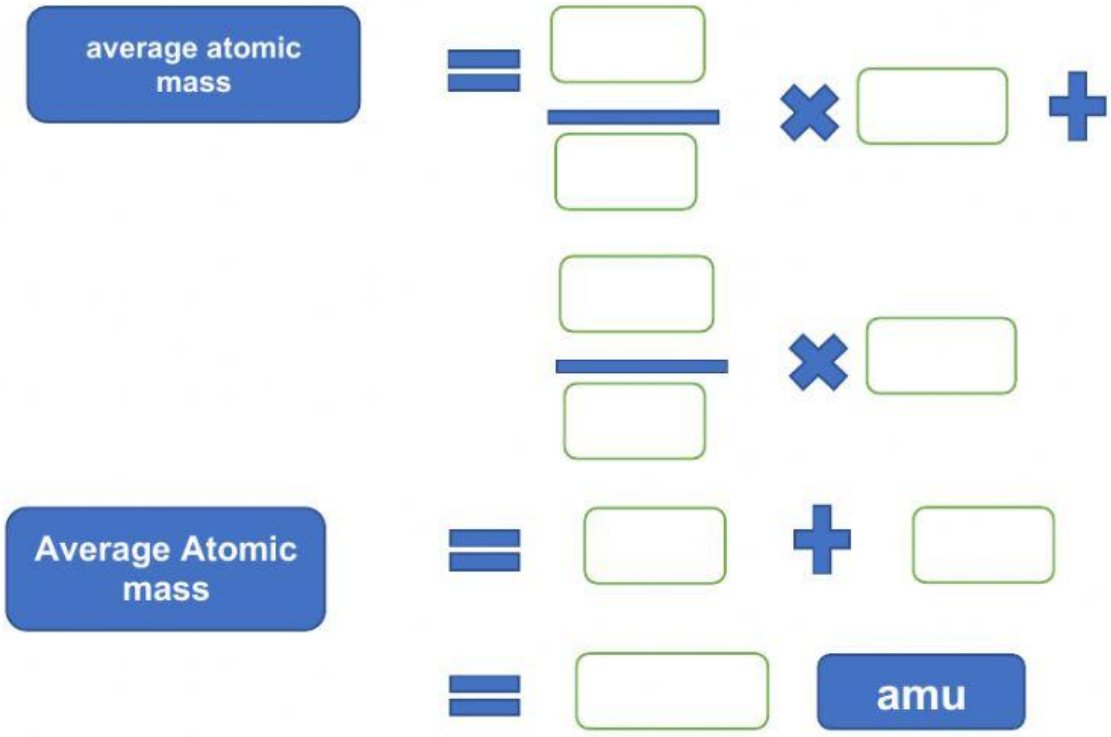
$\frac{1.11}{100} \times 13$

= 11.8668 + 0.1443

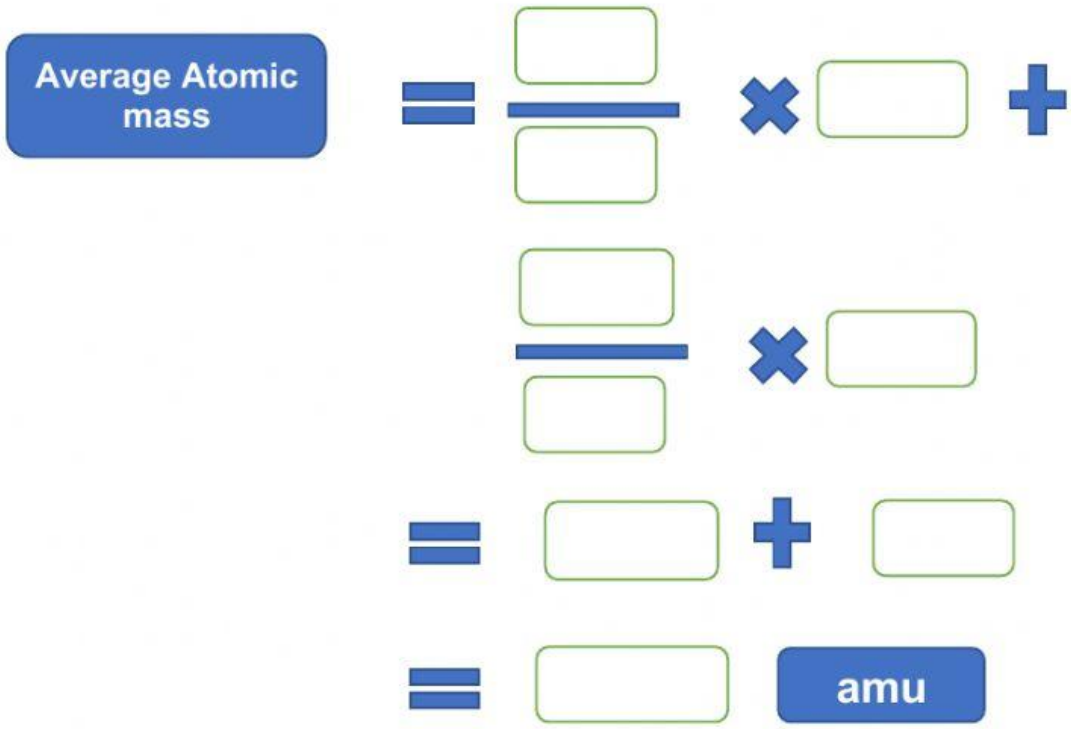
= 12.0111 amu

Calculate the average atomic mass of below isotopes

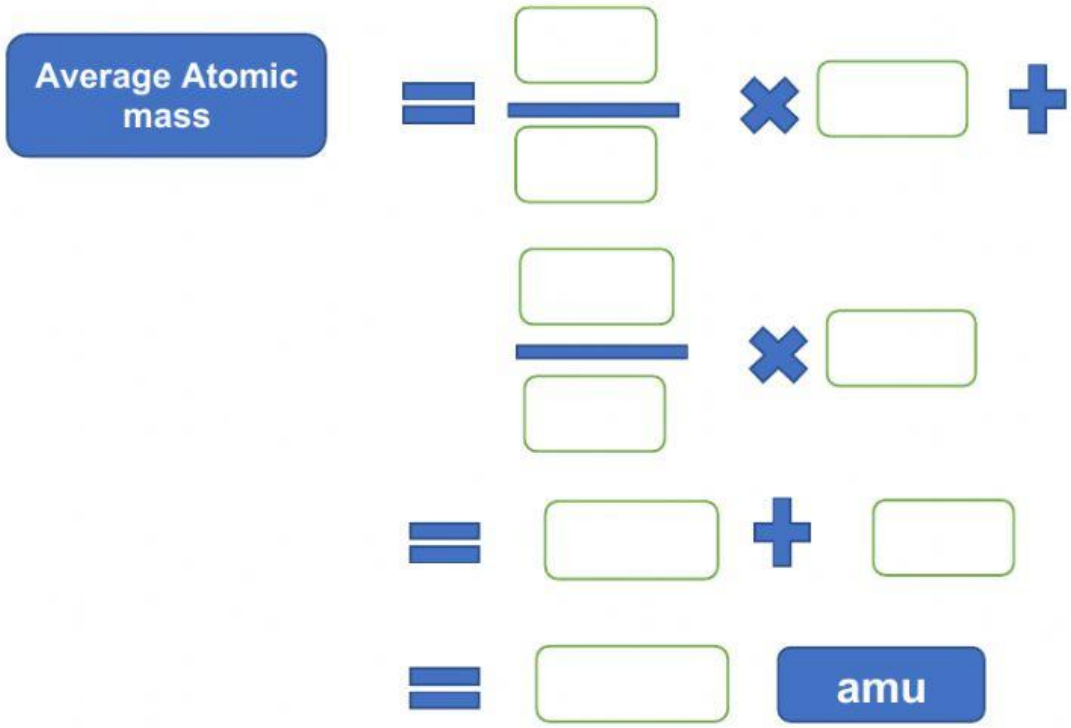
Isotope	% Abundance	Isotope Mass
Oxygen-16	99.76%	16
Oxygen-18	0.200%	18



Magnesium (Mg) Isotope	Percent Found in Nature
Mg-24	78.9%
Mg-25	10.0%



Isotopes	% Abundance	Mass in amu
Ga-69	60.11%	69
Ga-71	39.89%	71



Isotopes	% Abundance	Mass in amu
Copper-63	69.1%	62.9296
Copper-65	30.83%	64.9278