

Sitat of electrolyte and non-electrolyte

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 **LIVEWORKSHEETS**

Digital student worksheets (LKPD)

student name. :

Class :

No. Absence. :

Instruction:

Read and understand the following material:

Electrical conductivity in solution

A solution is a homogeneous mixture of two or more substances solution. Composed of a solvent and a solute. Based on the electrical conductivity, the nature of the solution can be influenced by the type of substance dissolved in a solution. Substances that can dissolve in water are divided into electtolytes and non-electrolytes. This difference is based on the presence of electrical conductivity in the solution.

K-3 Understand, apply, and analyze factual, conceptual, and procedural knowledge, based on their curiosity about science,technology,art,culture, and humanities with insight into humanity, nationality, state, and civilization, related to the causes of phenomena and events in a specific field of work to solve the problem

K-4 rocessing, reasoning, and presenting in the concrete and abstract realms are related to the development of what they learn in school independently, and being able to carry out specific tasks under direct supervision.

KD	KD FORMULA
KD. 3.5	Analyzing the properties of electrolyte and non-electrolyte solutions
KD. 4.5	Distinguish between the properties of electrolyte and non-electrolyte solutions.

Competency Achievement indicators (IPK)

CODE IPK	IPK formulation
3.5.1	Analyzing the symptoms of electric current conduction in a solution
3.5.2	Distinguish between electrolyte and non-electrolyte solutions
3.5.3	Identify electrolytes and non-electrolytes
3.5.4	Classify electrolytes and non-electrolytes
3.5.5	Distinguish between electrolyte and non-electrolyte solution based on current conduction electricity.



Purpose:

Through the discovery learning model, students are to analyze the state of the solution based on its electrical conductivity with the character of curiosity, active intelligent, and honest and can develop critical thinking skills, communicate, collaborate creatively (4C) in distinguishing the electrical conductivity of various solutions through the design and execution of experiments.

1. Following solution are classified as electrolyte solution
 - A. Alcohol
 - B. On coal
 - C. Kitchen salt water
 - D. Urea
 - E. Glukosa
2. The following pairs of solution are classified as non- electrolytes...
 - A. Urea and drinking water
 - B. Urea and sugar
 - C. Acetic acid and ammonia
 - D. Table salt and sulfuric acid
 - E. Sea water and table salt
3. Name the following chemical formula!

H₂SO₄

NaOH

4. Find a matching pair by drawing a line!

Ammonium
hydroxide

H_2SO_4

Sulfuric acid

NaCl

Sodium chloride

NH_4OH

