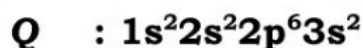
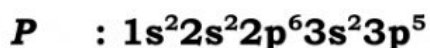
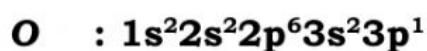


Question 4:

Electronic configuration for elements **O**, **P** and **Q** are as follows:



- a) State the period and group for **O** and **P** in the periodic table.

O: period = , **Group =**

P: period = , **Group =**

Q: period = , **Group =**

Example of hydride compound:

Boron hydride @ BH_3

- b) What is the molecular formula for chloride compound of **O**?

- c) What is the molecular formula hydride compound of **P**?

- d) Write the molecular formula of the compound formed when **O** reacts with **P**.

- e) Write the molecular formula of the compound formed when **Q** reacts with **P**.

- f) Which element is the biggest in term of size?
Which element has the highest IE1?

Size of atom

IE1

- g) Arrange in increasing order the first Ionisation energy of all the elements?

< <

Size of atom

Electronegativity

- h) Which element has the strongest electronegativity?