

- A. also be doubled
B. be halved
C. become one-fourth
D. remain the same
24. Which of the following stresses would lead the exothermic reaction below to shift to the right?
$$A(g) + B(g) \rightleftharpoons 3C(g) + D(aq)$$

A. Increasing [A]
B. Increasing [C]
C. Decreasing the volume
D. Increasing the temperature
25. Consider the following reaction system, which has a K_{eq} of 1.35×10^4 , taking place in a closed vessel at constant temperature.
$$AX_3(g) + X_2(g) \rightleftharpoons AX_5(g)$$

Which of the following is NOT true about this system at equilibrium?
A. Increasing the volume will produce more AX_3 .
B. AX_5 is the main compound present.
C. The rate of formation of AX_5 equals the rate of formation of AX_3 and X_2 .
D. Increasing the pressure will produce more AX_5 .
26. Ammonia gas form according to this equation;
$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$$

Experimental data shows that the reaction shifts to the left at very cold temperatures. Using this information, what type of reaction occur?
A. Exothermic
B. Maxwell Boltzmann-like
C. Endergonic
D. Endothermic
27. Which of the following does not affect the equilibrium of a reaction?
A. Adding heat
B. Increasing the concentration of reactants
C. Removing heat
D. Adding a catalyst
28. Consider the following reaction:
$$N_2O_4(g) \rightleftharpoons 2NO_2(g)$$

Which of the following changes would be expected to drive the reaction to the left?
A. An increase in pressure.
B. A decrease in pressure.
C. Addition of water to the reaction mixture.
D. Reducing the amount of NO_2 in the reaction vessel.
29. In the reaction:
$$2NO(g) + Cl_2(g) \rightleftharpoons 2NOCl(g)$$

If the pressure exerted by the equilibrium mixture of NO , Cl_2 , $NOCl$ gases is reduced at constant temperature, the system will re-establish equilibrium via
A. an addition of moles of Cl_2 .
B. an increase of the $NOCl$ concentration.
C. an addition of moles of $NOCl$.
D. an increase of equilibrium constant K_c .