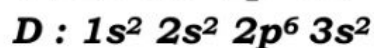
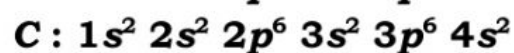
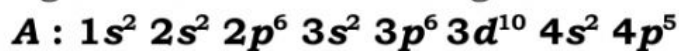
**Question 1:**

By referring to the electronic configuration of each element below:





(a) State the period, group and block for each element.

| Elements | Period | Block | Valence electron | Group | Ion      |
|----------|--------|-------|------------------|-------|----------|
| <b>A</b> |        |       |                  |       | <b>A</b> |
| <b>B</b> |        |       |                  |       | <b>B</b> |
| <b>C</b> |        |       |                  |       | <b>C</b> |
| <b>D</b> |        |       |                  |       | <b>D</b> |
| <b>E</b> |        |       |                  |       | <b>E</b> |

(b) State how the elements **A** to **E** are arranged in the periodic table?

The elements are arranged in the order of

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- (c) Why are elements **C** and **D** in the same group?  
**Both elements C and D have the same number of**  
\_\_\_\_\_.
- (d) Why are elements **A** and **C** in the same period?  
**Both elements A and C have the same**  
\_\_\_\_\_.
- (e) Between C and D, which is one of the elements is more electronegative?  
\_\_\_\_\_.
-   
Remember!  
*size of atom  $\propto \frac{1}{\text{electronegativity}}$*
- (f) Between A and C, which is one of the elements is more electronegative?  
\_\_\_\_\_.
- (g) Between C and D, which one of the elements has higher first ionization energy?  
\_\_\_\_\_.
-   
Remember!  
*size of atom  $\propto \frac{1}{\text{IE}_1}$*
- (h) Between A and C, which elements has higher first ionization energy (IE<sub>1</sub>)?  
\_\_\_\_\_.