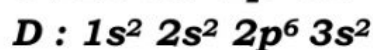
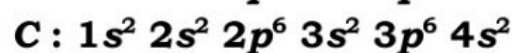
**Question 1:**

By referring to the electronic configuration of each element below:





(a) State the period, group and block for each element.

Elements	Period	Block	Valence electron	Group	Ion
A					A
B					B
C					C
D					D
E					E

(b) State how the elements **A** to **E** are arranged in the periodic table?

The elements are arranged in the order of

- (c) Why are elements **C** and **D** in the same group?
Both elements C and D have the same number of
_____.
- (d) Why are elements **A** and **C** in the same period?
Both elements A and C have the same
_____.
- (e) Between C and D, which is one of the elements is more electronegative?
_____.
- 
Remember!
- size of atom \propto electronegativity*
- (f) Between A and C, which is one of the elements is more electronegative?
_____.
- (g) Between C and D, which one of the elements has higher first ionization energy?
_____.
- 
Remember!
- size of atom \propto IE₁*
- (h) Between A and C, which elements has higher first ionization energy (IE₁)?
_____.