

1+						0																									
1						8																									
1	2+	Oxidation Number				8																									
1	2-	Valence Electrons				4																									
H Hydrogen 1.008	2	Family				He Helium 4.003																									
3	4					10																									
Li Lithium 6.941	Be Beryllium 9.0122	<p align="center">Some Polyatomic Ions and their Oxidation Numbers</p> <table border="0"> <tr> <td>1+</td> <td>1-</td> <td>2-</td> <td>3-</td> </tr> <tr> <td>ammonium (NH₄)</td> <td>acetate (C₂H₃O₂)</td> <td>carbonate (CO₃)</td> <td>phosphate (PO₄)</td> </tr> <tr> <td></td> <td>chlorate (ClO₃)</td> <td>sulfate (SO₄)</td> <td></td> </tr> <tr> <td></td> <td>hydroxide (OH)</td> <td></td> <td></td> </tr> <tr> <td></td> <td>nitrate (NO₃)</td> <td></td> <td></td> </tr> <tr> <td></td> <td>bicarbonate(HNO₃)</td> <td></td> <td></td> </tr> </table>				1+	1-	2-	3-	ammonium (NH ₄)	acetate (C ₂ H ₃ O ₂)	carbonate (CO ₃)	phosphate (PO ₄)		chlorate (ClO ₃)	sulfate (SO ₄)			hydroxide (OH)				nitrate (NO ₃)				bicarbonate(HNO ₃)				
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5	6					18																									
B Boron 10.81	C Carbon 12.011	N Nitrogen 14.007	O Oxygen 15.999	F Fluorine 18.998		Ne Neon 20.180																									
13	14	15	16	17		18																									
Al Aluminum 26.98	Si Silicon 28.086	P Phosphorus 30.974	S Sulfur 32.06	Cl Chlorine 35.453		Ar Argon 39.948																									
31	32	33	34	35		36																									
Ga Gallium 69.72	Ge Germanium 72.61	As Arsenic 74.922	Se Selenium 78.96	Br Bromine 79.904		Kr Krypton 83.80																									

	Symbols & Oxidation #	Formula	Name of Compound
15. Magnesium and chlorine			
16. Potassium and acetate			
17. Aluminum and sulfate			
18. Lithium and nitrogen			

19. An ion is an atom or group of atoms that has become electrically _____

20. When an atom loses an electron its charge is (*positive or negative*)

21. An ionic bond is the attraction between (*opposites, positive, neutral, or negative*) ions.

22. Ionic compounds are electrically (*charged, positive, neutral, or negative*).

23. The sum of the charges for an ionic compound is _____.

The two answers must be in the right order.

24. An ionic compound is the result of the bonding of a (*non-metal, metalloid, metal, noble gas*) with a (*non-metal, metalloid, metal*).