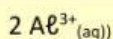
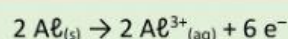
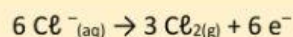
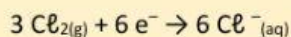
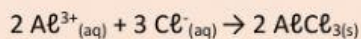
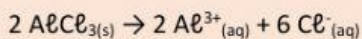


ELETRÓLISE: Reação de oxirredução não espontânea

ELETRÓLISE ÍGNEA

Agora é sua vez 1: Encontre a reação global para a eletrólise ígnea do Cloreto de Alumínio (AlCl_3).



Relacione as cores

Reação de Dissolução do Sal:

Semirreação do Cátodo:

Semirreação do Ânodo

Reação Global: _____ $\rightarrow$ _____ + _____

ELETRÓLISE AQUOSA

No cátodo, polo -

demais cátions



cátions da Família 1A, 2A e Al^{3+}

No ânodo, polo +

Ânions não-oxigenados



ânions oxigenados e F^{-}

Prioridade da Descarga

Bom Estudo!!!

Profa. Ana Paula Ruas (QuimicAna)



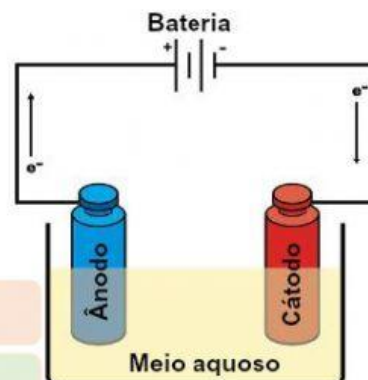
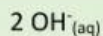
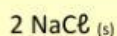
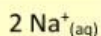
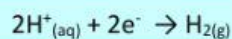
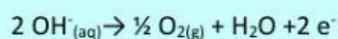
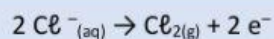
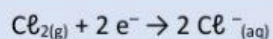
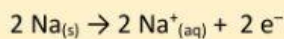
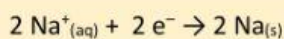
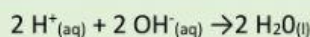
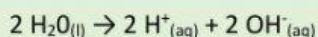
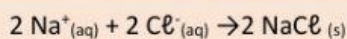
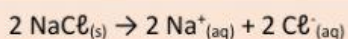
Fique atento:

Autoionização da água: $\text{H}_2\text{O}_{(l)} \rightarrow \text{H}^+_{(aq)} + \text{OH}^-_{(aq)}$

Descarga dos íons H^+ no cátodo: $2 \text{H}^+_{(aq)} + 2 \text{e}^- \rightarrow \text{H}_{2(g)}$

Descarga dos íons OH^- no ânodo: $2 \text{OH}^-_{(aq)} \rightarrow \frac{1}{2} \text{O}_{2(g)} + \text{H}_2\text{O}_{(l)} + 2 \text{e}^-$

Agora é sua vez 2: Eletrólise aquosa do NaCl em solução aquosa com eletrodos inertes



Célula eletrolítica

Relacione as cores

Reação de Dissolução do Sal:

Autoionização da água:

Semirreação do Cátodo:

Semirreação do Ânodo

Reação Global:



Bom Estudo!!!

Profa. Ana Paula Ruas (QuimicAna)

