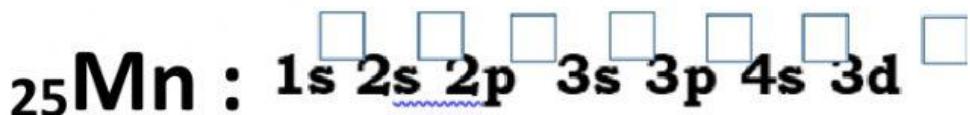


Exercise Electronic configuration of Manganese

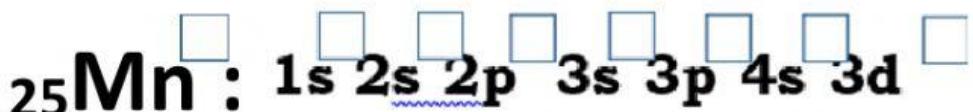
1) Write the electronic configuration of Manganese at its ground state:



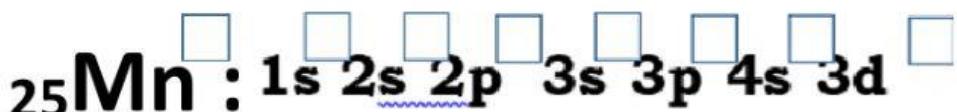
2) How many valence e belong to the Manganese:

3) Determine the outermost orbital in Manganese:

4) Write the electronic configuration of Mn ion after 2 e are donated:



5) Write the electronic configuration of Mn ion after 4 e are donated:



Give a set of quantum numbers for the e in the lowest energy orbital:

1st e: n = , l = , m= , s=

or

2nd e: n = , l = , m= , s=

