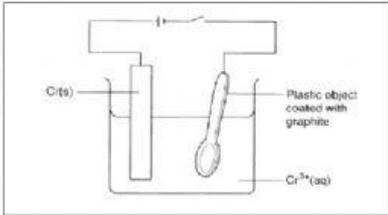
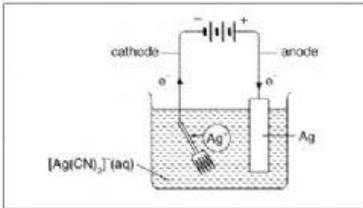


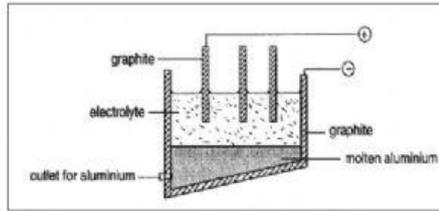
Instruction: Fill in the blanks for the materials of electrodes in each applications



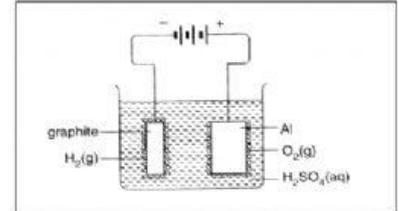
Electrode	Material	Equation
Anode		$\text{Cr} \rightarrow \text{Cr}^{3+} + 3\text{e}^-$
Cathode		$\text{Cr}^{3+} + 3\text{e}^- \rightarrow \text{Cr}$
Electrolyte	Chromium(III) salt	



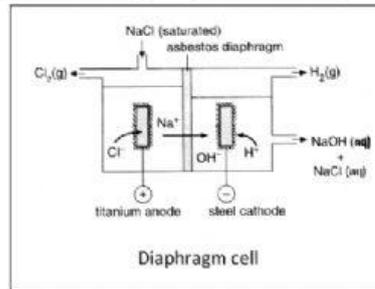
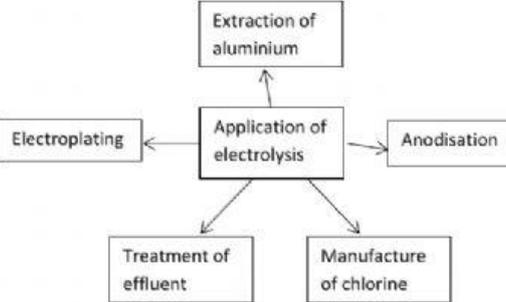
Electrode	Material	Equation
Anode		$\text{Ag} \rightarrow \text{Ag}^+ + \text{e}^-$
Cathode		$\text{Ag}^+ + \text{e}^- \rightarrow \text{Ag}$
Electrolyte	[Ag(CN)₂]⁻	



Electrode	Material	Equation
Anode		$2\text{O}^{2-} \rightarrow \text{O}_2 + 4\text{e}^-$
Cathode		$\text{Al}^{3+} + 3\text{e}^- \rightarrow \text{Al}$
Electrolyte	Al ₂ O ₃ (Bauxite)	



Electrode	Material	Equation
Anode		$2\text{H}_2\text{O} \rightarrow \text{O}_2 + 4\text{H}^+ + 4\text{e}^-$ $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
Cathode		$2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$
Electrolyte	H ₂ SO ₄	



Electrode	Material	Equation
Anode		$2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$
Cathode		$2\text{H}_2\text{O} + 2\text{e}^- \rightarrow \text{H}_2 + 2\text{OH}^-$
Electrolyte	Brine, NaCl(aq)	