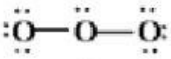
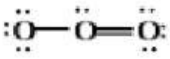
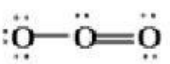
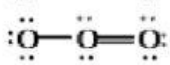
CHEMBUDDY CHAPTER 4
4.1 LEWIS STRUCTURE

CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	Identify the species that NOT involves dative bond. A. H_2O B. H_3O^+ C. NH_4^+ D. Al_2Cl_6	A B C D
2	How many resonance structures can be drawn for phosphate ions, PO_4^{3-}? A. 2 C. 4 B. 3 D. none	A B C D
3	The most plausible Lewis structure for a chlorate ion, ClO_3^-, should show contain ____ single bond(s), ____ double bond(s), and ____ lone pair(s). A. 2, 1, 10 C. 3, 0, 10 B. 1, 2, 8 D. 2, 1, 9	A B C D
4	Which Lewis structure best represents ozone molecule, O_3? A.  B.  C.  D. 	A B C D

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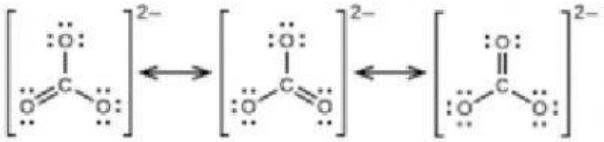
5	The formal charge on the bromine atom in BrO_3^- drawn with three single bonds is A. -2 C. +1 B. -1 D. +2	A B C D
6	11. What is the picture showing?  A. The resonance structure of carbonate ions. B. The formal charge of each atom in carbonate. C. The electronegativity of carbonate. D. The geometrical shape of carbonate.	A B C D
7	Which molecule has central atom that obeys octet rule? A. SF_4 C. OF_2 B. BCl_3 D. $SnCl_2$	A B C D
8	What element can form an incomplete octet? A. Boron C. Bismuth B. Bromine D. Barium	A B C D

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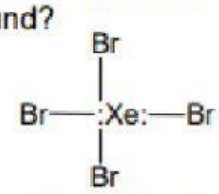
9	The electronic configuration of elements X and Y are as follows: X: $1s^2\ 2s^2\ 2p^1$ Y: $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^5$ When X and Y combine, the most plausible formula and bonding are FORMULA BONDING A. XY_3 covalent B. X_2Y covalent C. XY_2 ionic D. X_2Y_3 ionic	A B C D
10	What is the unusual feature about this compound?  A. Obey the octet rule B. Odd number electron C. Incomplete octet D. Expended octet	A B C D
11	The ion has no valence electron in its Lewis symbol is **(C1&C2)** A. B^{3+} C. Si^{+} B. O^{2-} D. Cl^{-}	A B C D

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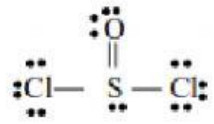
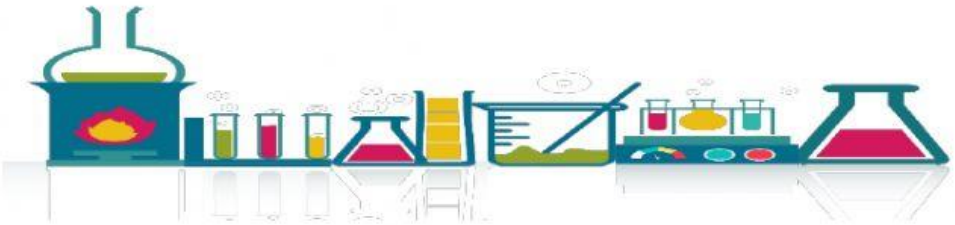
12	Lewis structure of $SOCl_2$ is given as follows **(C1&C2)**  The formal charges on sulphur, chlorine and oxygen are A. Sulphur 0, Chlorine 0, Oxygen 0 B. Sulphur -1, Chlorine +2, Oxygen -1 C. Sulphur +2, Chlorine -1, Oxygen -1 D. Sulphur +2, Chlorine -2, Oxygen 0	A B C D
13	Which of the following is NOT the case of exceptional to octet rule: A. OF_2 B. ClO_2^+ C. $SnCl_2$ D. $BeCl_2$	A B C D



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