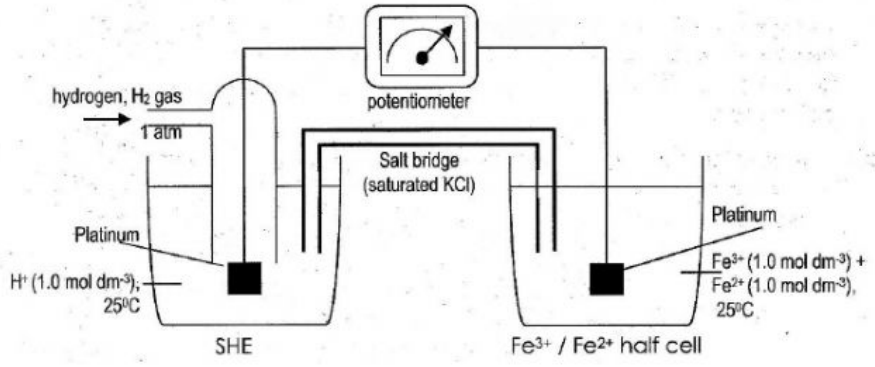
Fill in the blanks



E^o for H^+/H_2 = 0V, E^o for Fe^{3+}/Fe^{2+} = +0.77V

Oxidation occur at and reduction occur at the . During the redox reaction, H_2 is to H^+ and the oxidation number from 0 to +1. While Fe^{3+} is to Fe^{2+} and the oxidation number from +3 to +2. The oxidising agent is and the reducing agent is .

The observation is solution turn into solution

decrease	increase	H_2	oxidised	reduced
Fe^{3+}/Fe^{2+}half cell	green	Fe^{3+}	SHE	yellow