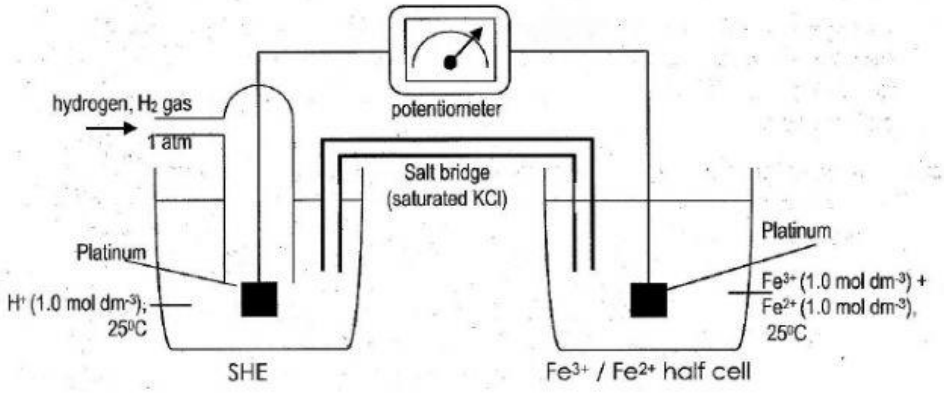
Fill in the blanks



E^o for H^+/H_2 = 0V, E^o for Fe^{3+}/Fe^{2+} = +0.77V

The half-equation at anode is →

The half-equation at cathode is →

The overall/ionic equation for the cell is

→

The E_{cell} of the cell is

The cell diagram for the cell is

||

The electron flow from (SHE// Fe^{3+}/Fe^{2+}) to (SHE//Fe^{3+}/Fe^{2+}) half cell through external circuit.

$Fe^{2+}(aq)$	$H_2(g)$
$Fe^{3+}(aq)$, $Fe^{2+}(aq)$ | $Pt(s)$	$2H^+(aq) + 2e$
$Fe^{3+}(aq) + e$	$2H^+(aq) + 2Fe^{2+}(aq)$
$H_2(g) + 2Fe^{3+}(aq)$	$Pt(s) \| H_2(g) \| H^+(aq)$

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