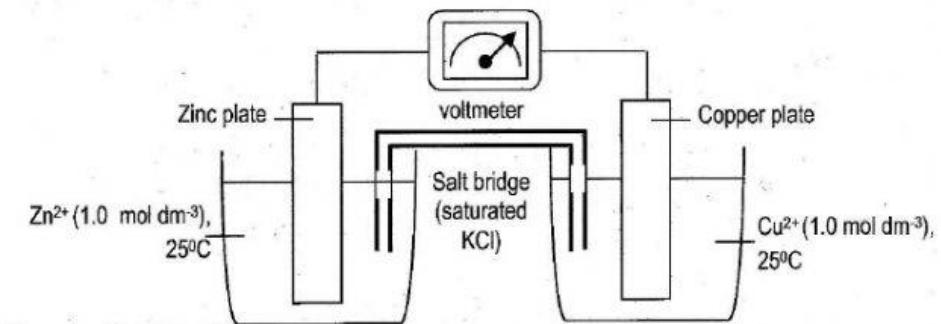


### Fill in the blanks



$E^\circ$  for  $\text{Zn}^{2+}/\text{Zn} = -0.76\text{V}$ ,  $E^\circ$  for  $\text{Cu}^{2+}/\text{Cu} = +0.34\text{V}$

Oxidation occurs at the anode and reduction occurs at the cathode. During the redox reaction,  $\text{Fe}^{2+}$  is oxidised to  $\text{Fe}^{3+}$  and the oxidation number increases from 0 to +2. While  $\text{Cu}^{2+}$  is reduced to  $\text{Cu}^{+}$  and the oxidation number decreases from +2 to 0. The oxidising agent is  $\text{Fe}^{3+}$  and the reducing agent is  $\text{Cu}^{2+}$ .

The observations are:

- plate dissolves/corroded
- The solution of copper(II) salt faded and brown solid deposited