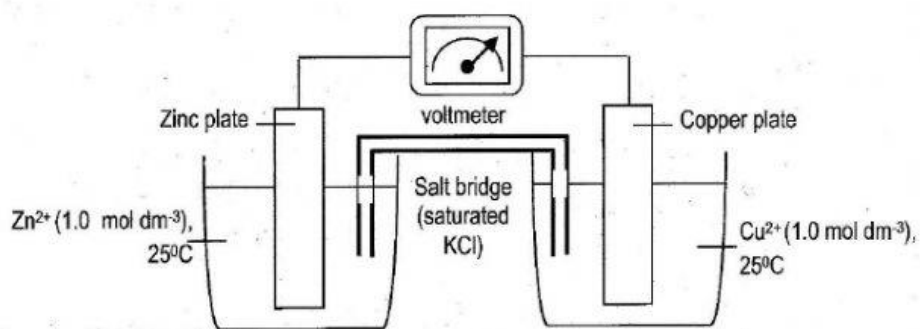


Fill in the blanks



$$E^\circ \text{ for } \text{Zn}^{2+}/\text{Zn} = -0.76\text{V}, E^\circ \text{ for } \text{Cu}^{2+}/\text{Cu} = +0.34\text{V}$$

Oxidation occur at \_\_\_\_\_ plate and reduction occur at the \_\_\_\_\_ plate. During the redox reaction, \_\_\_\_\_ is oxidised to \_\_\_\_\_ and the oxidation number increase from 0 to +2.

While \_\_\_\_\_ is reduced to \_\_\_\_\_ and the oxidation number decrease from +2 to 0. The oxidising agent is \_\_\_\_\_ and the reducing agent is \_\_\_\_\_

The observations are:

- \_\_\_\_\_ plate dissolves/corroded
- The \_\_\_\_\_ solution of copper(II) salt faded and brown solid deposited