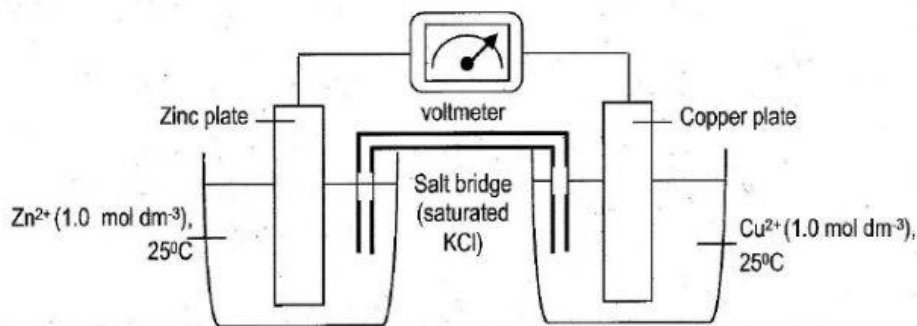


Fill in the blanks based on the diagram below:



$$E^\circ \text{ for } \text{Zn}^{2+}/\text{Zn} = -0.76\text{V}, E^\circ \text{ for } \text{Cu}^{2+}/\text{Cu} = +0.34\text{V}$$

The half-equation at anode is

→

The half-equation at cathode is

→

The overall/ionic equation for the cell is

→

The  $E_{\text{cell}}$  of the cell is

The cell diagram for the cell is

||

The electron flow from

plate to

plate through external circuit

<b>Cu(s)</b>	<b>zinc</b>
<b>copper</b>	<b>Zn(s)</b>
<b><math>\text{Cu}^{2+}(\text{aq}) \mid \text{Cu}(\text{s})</math></b>	<b><math>\text{Zn}(\text{s}) + \text{Cu}^{2+}(\text{aq})</math></b>
<b><math>\text{Zn}^{2+}(\text{aq}) + \text{Cu}(\text{s})</math></b>	<b><math>\text{Zn}(\text{s}) \mid \text{Zn}^{2+}(\text{aq})</math></b>
<b><math>\text{Cu}^{2+}(\text{aq}) + 2\text{e}^-</math></b>	<b><math>\text{Zn}^{2+}(\text{aq}) + 2\text{e}^-</math></b>