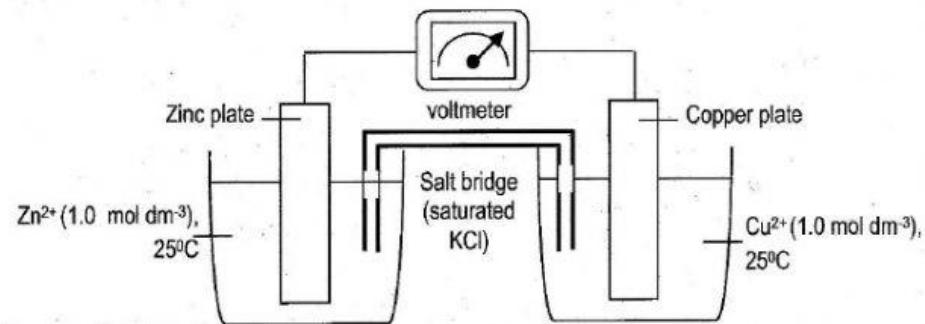


Fill in the blanks based on the diagram below:



$$E^\circ \text{ for } \text{Zn}^{2+}/\text{Zn} = -0.76\text{V}, E^\circ \text{ for } \text{Cu}^{2+}/\text{Cu} = +0.34\text{V}$$

The half-equation at anode is \rightarrow

The half-equation at cathode is \rightarrow

The overall/ionic equation for the cell is \rightarrow

The E_{cell} of the cell is

The cell diagram for the cell is ||

The electron flow from plate to plate through external circuit

Cu(s)	zinc
copper	Zn(s)
Cu²⁺(aq) Cu(s)	Zn(s) + Cu²⁺(aq)
Zn²⁺(aq) + Cu(s)	Zn(s) Zn²⁺(aq)
Cu²⁺(aq) + 2e	Zn²⁺(aq) + 2e