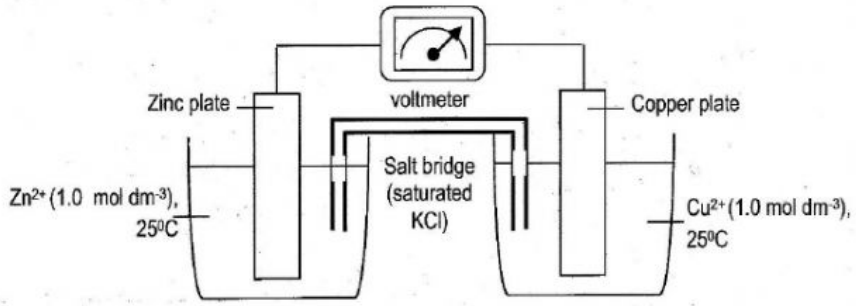
Fill in the blanks based on the diagram below:



E^{o} for Zn^{2+}/Zn = -0.76V, E^{o} for Cu^{2+}/Cu = +0.34V

The half-equation at anode is →

The half-equation at cathode is →

The overall/ionic equation for the cell is →

The E_{cell} of the cell is

The cell diagram for the cell is ||

The electron flow from plate to plate through external circuit

Cu(s)	**zinc**
copper	**Zn(s)**
$Cu^{2+}(aq)$ | Cu(s)	**$Zn(s) + Cu^{2+}(aq)$**
$Zn^{2+}(aq) + Cu(s)$	**Zn(s) | $Zn^{2+}(aq)$**
$Cu^{2+}(aq) + 2e$	**$Zn^{2+}(aq) + 2e$**