

CHEMBUDDY CHAPTER 3

3.1 CLASSIFICATION OF ELEMENTS



CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	What is group in periodic table? A. The horizontal row with label 1 to 7 B. The horizontal row with label 1 to 18 C. The vertical columns with label 1 to 18 D. The vertical columns with label 1 to 7 only	A B C D
2	What is period in periodic table? A. The horizontal row with label 1 to 7 B. The horizontal row with label 1 to 18 C. The vertical columns with label 1 to 18 D. The vertical columns with label 1 to 7 only	A B C D
3	Which block is NOT TRUE in periodic table? A. b block B. d block C. f block D. p block	A B C D
4	Deduce period of element Y in periodic table. Given the electronic configuration of Y as $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$ A. period 1 B. period 2 C. period 3 D. period 4	A B C D
5	What is the charge on all the simple ions of elements of group 17? A. $1+$ B. $2+$ C. $1-$ D. $2-$	A B C D
6	Sodium ($Z=11$) and Silicon ($Z=14$) are examples of two elements which belong to the same A. group B. period C. block D. compound	A B C D

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	$^{37}_{17}P$ What is the position of this element in the periodic table?																										
7	<table> <thead> <tr> <th></th> <th>Block</th> <th>Group</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>p</td> <td>17</td> <td>3</td> <td>A</td> </tr> <tr> <td>B.</td> <td>p</td> <td>15</td> <td>3</td> <td>B</td> </tr> <tr> <td>C.</td> <td>p</td> <td>7</td> <td>3</td> <td>C</td> </tr> <tr> <td>D.</td> <td>s</td> <td>5</td> <td>3</td> <td>D</td> </tr> </tbody> </table>		Block	Group	Period		A.	p	17	3	A	B.	p	15	3	B	C.	p	7	3	C	D.	s	5	3	D	
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8	$^{63}_{29}Q$ What is the position of this element in the periodic table? <table> <thead> <tr> <th></th> <th>Block</th> <th>Group</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>s</td> <td>2</td> <td>4</td> <td>A</td> </tr> <tr> <td>B.</td> <td>s</td> <td>12</td> <td>4</td> <td>B</td> </tr> <tr> <td>C.</td> <td>d</td> <td>9</td> <td>3</td> <td>C</td> </tr> <tr> <td>D.</td> <td>d</td> <td>11</td> <td>4</td> <td>D</td> </tr> </tbody> </table>		Block	Group	Period		A.	s	2	4	A	B.	s	12	4	B	C.	d	9	3	C	D.	d	11	4	D	
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9	An element R has a valence electronic configuration of $5s^2 4d^6$. What is the period and block of R in the Periodic Table? <table> <thead> <tr> <th></th> <th>Block</th> <th>Period</th> <th></th> </tr> </thead> <tbody> <tr> <td>A.</td> <td>s</td> <td>5</td> <td>A</td> </tr> <tr> <td>B.</td> <td>s</td> <td>4</td> <td>B</td> </tr> <tr> <td>C.</td> <td>d</td> <td>5</td> <td>C</td> </tr> <tr> <td>D.</td> <td>d</td> <td>4</td> <td>D</td> </tr> </tbody> </table>		Block	Period		A.	s	5	A	B.	s	4	B	C.	d	5	C	D.	d	4	D						
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10	An electronic configuration for element T is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$. Determine the group of element T A. 4 B. 14 C. 6 D. 16																										
11	An electronic configuration for element U is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^8$. Determine the group of element U. A. 8 B. 10 C. 14 D. 16																										

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	The proton number for element V is 32. What is the valence orbital of element V? A. 3d 4p B. 4p C. 4s 4p D. 3d	A B C D
12	Elements in the Periodic Table are arranged in order increasing A. nucleon number B. proton number C. number of electron D. number of neutron	A B C D
13	Elements in the same group will have similar A. conductor properties B. thermal properties C. physical properties D. chemical properties	A B C D
14	Elements in Group 2 are known as A. alkaline earth metal B. noble gases C. alkali metal D. halogen	A B C D
15		

