

CHEMBUDDY CHAPTER 3

3.1 CLASSIFICATION OF ELEMENTS



CHOOSE THE CORRECT ANSWER

NO	QUESTION	ANSWER
1	<p>What is group in periodic table? (C1&C2)</p> <p>A. The horizontal row with label 1 to 7</p> <p>B. The horizontal row with label 1 to 18</p> <p>C. The vertical columns with label 1 to 18</p> <p>D. The vertical columns with label 1 to 7 only</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>
2	<p>What is period in periodic table? (C1&C2)</p> <p>A. The horizontal row with label 1 to 7</p> <p>B. The horizontal row with label 1 to 18</p> <p>C. The vertical columns with label 1 to 18</p> <p>D. The vertical columns with label 1 to 7 only</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>
3	<p>Which block is NOT TRUE in periodic table? (C1&C2)</p> <p>A. b block</p> <p>B. d block</p> <p>C. f block</p> <p>D. p block</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>
4	<p>Deduce period of element X in periodic table. Given the electronic configuration of Y as $Y: 1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$ (C1&C2)</p> <p>A. Period 1</p> <p>B. Period 2</p> <p>C. Period 3</p> <p>D. Period 4</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>
5	<p>What is the charge on all the simple ions of elements of group 17?</p> <p>A. 1+</p> <p>B. 2+</p> <p>C. 1-</p> <p>D. 2-</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>
6	<p>Sodium (Z=11) and Silicon (Z=14) are examples of two elements which belong to the same</p> <p>A. group</p> <p>B. period</p> <p>C. block</p> <p>D. compound</p>	<p>A</p> <p>B</p> <p>C</p> <p>D</p>



7	<div>$^{37}_{17}\text{P}$</div> <p>What is the position of this element in the periodic table?</p> <table><thead><tr><th></th><th>Block</th><th>Group</th><th>Period</th></tr></thead><tbody><tr><td>A.</td><td>p</td><td>17</td><td>3</td></tr><tr><td>B.</td><td>p</td><td>15</td><td>3</td></tr><tr><td>C.</td><td>p</td><td>7</td><td>3</td></tr><tr><td>D.</td><td>s</td><td>5</td><td>3</td></tr></tbody></table>		Block	Group	Period	A.	p	17	3	B.	p	15	3	C.	p	7	3	D.	s	5	3	A B C D
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8	<div>$^{63}_{29}\text{Q}$</div> <p>What is the position of this element in the periodic table?</p> <table><thead><tr><th></th><th>Block</th><th>Group</th><th>Period</th></tr></thead><tbody><tr><td>A.</td><td>s</td><td>2</td><td>4</td></tr><tr><td>B.</td><td>s</td><td>12</td><td>4</td></tr><tr><td>C.</td><td>d</td><td>9</td><td>3</td></tr><tr><td>D.</td><td>d</td><td>11</td><td>4</td></tr></tbody></table>		Block	Group	Period	A.	s	2	4	B.	s	12	4	C.	d	9	3	D.	d	11	4	A B C D
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9	<p>An element R has a valence electronic configuration of $5s^2 4d^6$. What is the period and block of X in the Periodic Table?</p> <table><thead><tr><th></th><th>Block</th><th>Period</th></tr></thead><tbody><tr><td>A.</td><td>s</td><td>5</td></tr><tr><td>B.</td><td>s</td><td>4</td></tr><tr><td>C.</td><td>d</td><td>5</td></tr><tr><td>D.</td><td>d</td><td>4</td></tr></tbody></table>		Block	Period	A.	s	5	B.	s	4	C.	d	5	D.	d	4	A B C D					
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10	<p>An electronic configuration for element T is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$. Determine the group of element Y</p> <p>A. 4 B. 14 C. 6 D. 16</p>	A B C D																				
11	<p>An electronic configuration for element U is $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^8$. Determine the group of element U.</p> <p>A. 8 B. 10 C. 14 D. 16</p>	A B C D																				

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12	The proton number for element V is 32. What is the valence orbital of element V? A. 3d 4p B. 4p C. 4s 4p D. 3d	A B C D
13	Elements in the Periodic Table are arranged in order increasing A. nucleon number B. proton number C. number of electron D. number of neutron	A B C D
14	Elements in the same group will have similar A. conductor properties B. thermal properties C. physical properties D. chemical properties	A B C D
15	Elements in Group 2 are known as A. alkaline earth metal B. noble gases C. alkali metal D. halogen	A B C D

