C. $PH_3 > H_2S > SiH_4$

D. $SiH_4 > PH_3 > H_2S$

19. Which of the following molecules forms trigonal bipyramidal?

A. PCl_3 C. NBr_3

B. PCl_5 D. ICl_4^+

20. Which of the following statement is **TRUE** for ICl_4^+ and ICl_4^-.

A. Both compounds form non-polar molecules.

B. There are more lone pairs on the central atom, I for ICl_4^+ than ICl_4^-.

C. Bond angle between Cl-I-Cl in ICl_4^+ is smaller than ICl_4^-.

D. Both compounds have 5 electron pairs surrounding the central atom.

21. PCl_5 molecule has

A. to be non polar molecule with non polar bonds.

B. non polar bonds and is a polar molecule.

C. to be polar molecule with polar bonds.

D. polar bonds but is a non polar molecule.

22. Choose the **CORRECT** statement about ammonia molecule.

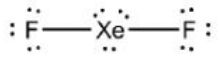
A. The molecular geometry is trigonal pyramidal.

B. The N-H bond is polar but the molecule is non-polar.

C. The bond angles between H-N-H is 109°.

D. It is not obey the octet rule.

23. Why XeF_2 is a non polar molecule?



A. Because it is expended octet molecule.

B. Because it has same terminal atoms.

C. Because the linear shape is symmetrical thus μ=0.

D. Because it has 3 lone pairs.

24. These molecules are sp^2 hybrid orbital on central atom except

A. CO_2 C. $BeCl_2$

B. H_2O D. HCN

25. What are the orbitals overlapped to form BF_3?

A. sp and p-orbital

B. sp^2 and p-orbital

C. sp^3 and p-orbital

D. sp^3 and s-orbital

26. The number of π and σ bond C_2H_4 are

A. 1 π and 4 σ

B. 1 π and 6 σ

C. 2 π and 4 σ

D. 1 π and 5 σ