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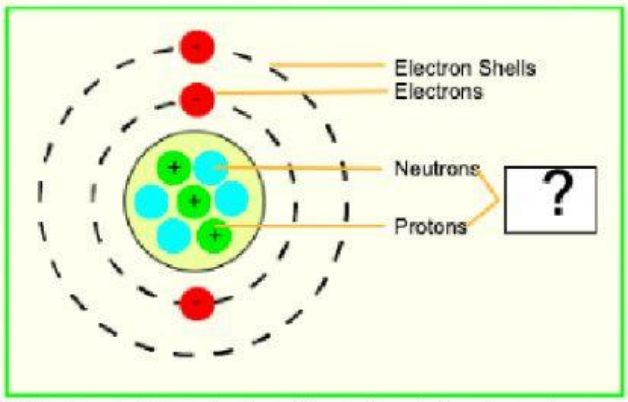
FORMATIVE TEST 1

A. Multiple-choice Questions

1. Look at the picture.



Which label is missing from the following diagram?
A. Neutroneus
B. Proton
C. Necleus
D. Electron

2. The number of protons in an atom always is the same as the number of …
A. Electron
B. Neutron
C. Nuclei
D. Atomic number

3. The maximum mass of an atom is concentrated in which of these?
A. Nucleus
B. Neutrons
C. Protons
D. Electrons

4. An atom has a mass number of 37 and atomic number 17. How many neutrons does it have?
A. 20
B. 54
C. 17
D. 21

B. Fill the questions with the correct answers.

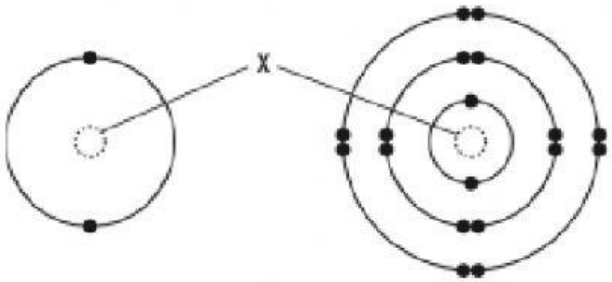
1. Vanadium has two isotopes.

$^{50}_{23}V$ $^{51}_{23}V$

An atom contains protons, electrons and neutrons.
Complete the table to show the number of protons, electrons and neutrons in these two isotopes of vanadium.

Isotope	Number of protons	Number of electrons	Number of neutrons
$^{50}_{23}V$	23	23	
$^{51}_{23}V$			28

2. Helium and argon are noble gases. The atomic structures of helium and argon are shown below.



These are 2 statements are correct. Give a tick for the correct statement.

Argon has an complete inner shell of electrons.

Helium has a complete outer shell of electrons.

Helium has an incomplete outer shell of electrons.

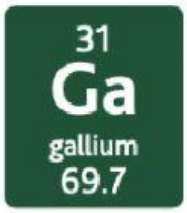
An atom of argon has 16 electrons.

3. Complete these sentences.

   (i) ..... of an element are the atoms of same element having same atomic number but different mass number.

   (ii) The atoms $_{18}Ar^{40}$ and $_{19}K^{40}$ are pairs of .....

4. Tick the correct electron configuration from these atom.



$1s^2\ 2s^2\ 3s^2\ 3p^6\ 3d^{10}\ 4s^2\ 4p^1$

$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 4d^{10}\ 4p^1$

$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}\ 4p^1$

$1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^2\ 3d^{10}\ 4d^1$