

Quiz No. 1
Electron Configuration

I. Do the following.

- A. Write the complete *electron configuration* for each atom.
- B. Give the Highest Energy Level of the electron
- C. Give the valence electron
- D. Write your answers inside the box.

NOTE: Please *follow the format* when you write the **electron configuration** to get the correct answer.



Example:
Neon = $1s^2 2s^2 2p^6$



1. Argon **Electron Configuration –**

Highest Energy Level –
Valence Electron –

2. Sulfur **Electron Configuration –**

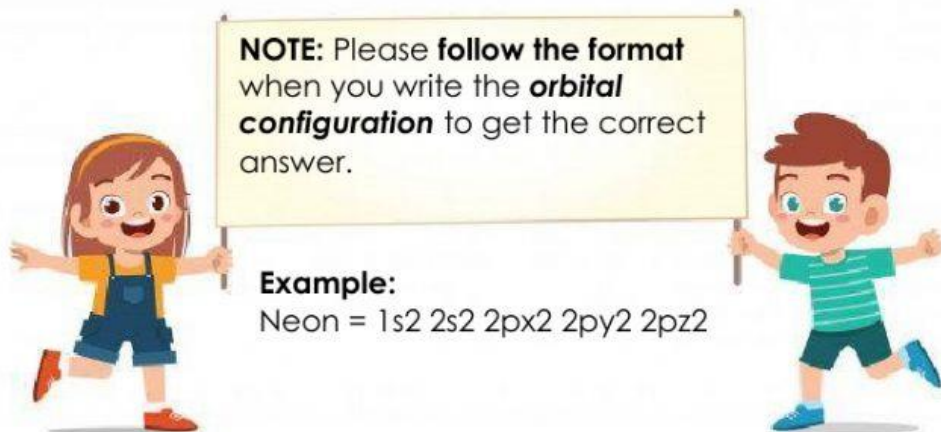
Highest Energy Level –
Valence Electron –

3. Radium **Electron Configuration –**

Highest Energy Level –
Valence Electron –

II. Identify what elements are represented by each of the following electron configurations? (4 pts)

- a. $1s^2 2s^2 2p^5$
- b. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- c. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^4$



III. Write the **orbital configuration** of the following elements.

a. Aluminum –

b. Selenium –