

1. CHEMICAL EQUATIONS

- Burning of crackers is a.... []
 A) physical change
 B) chemical change
 C) Chemical decomposition reaction
 D) None of these.
- Adding water to powdered lime is.... []
 A) Physical change
 B) Chemical change
 C) Endothermic reaction
 D) Chemical decomposition reaction
- Heating of lime stone gives carbon dioxide and powdered lime. This is a ...[]
 A) Chemical change
 B) Exothermic reaction
 C) Endothermic reaction
 D) Chemical combination reaction.
- When dilute hydrochloric acid is added to zinc pieces,... []
 A) It needs catalyst
 B) Precipitate is formed
 C) Heat is liberated
 D) Heat is absorbed
- $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$ In this chemical reaction... []
 A) Na_2SO_4 , BaCl_2 are reactants
 B) BaSO_4 , NaCl are products
 C) A white precipitate is formed.
 D) A and C.
- Which of the following is correct ? []
 A) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
 B) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 C) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
 D) All the above
- Which of the following is NOT correct,[]
 A) Calcium oxide + Water $\rightarrow$ Calcium Hydroxide
 B) Hydrogen + Oxygen $\rightarrow$ Water
 C) Calcium oxide $\rightarrow$ Water + Calcium hydroxide
 D) A and B
- Word equation for $\text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$ []
 A) Zinc + Hydrochloric acid $\rightarrow$ Zinc chloride + Hydrogen
 B) Zinc + Cupric chloride $\rightarrow$ Zinc chloride + Copper
 C) Zinc + Sulphuric acid $\rightarrow$ Zinc sulphate + Hydrogen
 D) Zinc + Copper chloride $\rightarrow$ Zinc chloride + Copper
- Choose the chemical equation for: Zinc + Hydrochloric acid $\rightarrow$ Zinc chloride + Hydrogen
 A) $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
 B) $\text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$
 C) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
 D) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- Choose the chemical equation for: Sodium sulphate + Barium chloride $\rightarrow$ Barium sulphate + Sodium chloride
 A) $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
 B) $\text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$
 C) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
 D) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- Choose the chemical equation for: Calcium carbonate $\rightarrow$ Calcium oxide + Carbon dioxide
 A) $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$
 B) $\text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$
 C) $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
 D) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
- write chemical equations for the following: Magnesium + Iodine $\rightarrow$ Magnesium Iodide
 A) $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$
 B) $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$
- write chemical equations for the following: Calcium hydroxide + Nitric acid $\rightarrow$ Water + Calcium nitrate
 A) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
 B) $\text{Ca(OH)}_2 + \text{HNO}_3 \rightarrow \text{H}_2\text{O} + \text{Ca(NO}_3)_2$
- write chemical equations for the following: Sodium Hydroxide + Sulphuric acid $\rightarrow$ Sodium sulphate + Water
 A) $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
 B) $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- write chemical equations for the following: Sodium + Water $\rightarrow$ Sodium Hydroxide + Hydrogen
 A) $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g})$
 B) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
- Formula unit of magnesium bromide is... []
 A) MgBr
 B) NaCl
 C) $\text{Mg}^{+2}, \text{Br}^-$
 D) Na^+, Cl^-
- $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$ coefficient required for NaCl for balancing equation is.... []
 A) 1
 B) 2
 C) 3
 D) 4
- Identify the balanced equation []
 A) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
 B) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 C) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
 D) A and C
- Which chemical equations are not balanced. []
 A) $\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$
 B) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
 C) $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 D) A and C