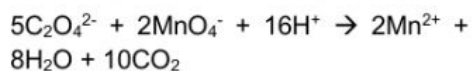


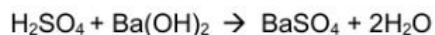
28. A 25 mL solution containing $\text{C}_2\text{O}_4^{2-}$ ions was titrated with KMnO_4 in acidic medium as follow:



Calculate the concentration of $\text{C}_2\text{O}_4^{2-}$ ions if the titration required 30.50 mL of 1.25 M KMnO_4 solution.

- A. 3.81 M C. 4.23 M
B. 3.32 M D. 4.78 M

29. The reaction between sulphuric acid, H_2SO_4 and barium hydroxide, $\text{Ba}(\text{OH})_2$ is shown below



Calculate the mass of $\text{Ba}(\text{OH})_2$ required to produce 4.35g BaSO_4 , if the percentage yield of the reaction is 85%.

- A. 3.75 g C. 3.48 g
B. 6.97 g D. 5.12 g