NAME: _____ DATE: _____

Let's start calculate the Molarity!

EXPERIMENT 2: DATA ANALYSIS

READ THE QUESTION FIRST.

A titration of 25.00 mL of an ***x*** M HCl solution with 0.15 M NaOH starts at a burette reading for NaOH of 0.20 mL. The burette reading of the end point is 24.10 mL.

How much is NaOH used in this exp?

Volume of NaOH dispensed

=

=

Now calculate the no. of moles of NaOH dispensed.

Moles of NaOH = MV

= ×

=

We need to know how many moles of NaOH so we can determine no.of moles of HCl reacted.

Let's write the balanced equation for the neutralisation reaction.

Balanced equarion:

+ → +

Calculate the number of HCl present in the acid solution. (DO stoichiometry)

react with

react with

Formula

Molarity = _____

Substitution & Final answer

= _____

x =