

In the space below, write the unabbreviated electron configurations of the following elements:

8
O
oxygen
16.0

11
Na
sodium
23.0

26
Fe
iron
55.8

35
Br
bromine
79.9

56
Ba
barium
137.3

7
N
nitrogen
14.0

17
Cl
chlorine
35.5

$1s^2 2s^2 2p^4$

$1s^2 2s^2 2p^6 3s^1$

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$

$1s^2 2s^2 2p^3$

$1s^2 2s^2 2p^6 3s^2 3p^5$

$1s^2 2s^2 2p^6 3s^2 3p^6$

Determine what elements are denoted by the following electron configurations:

14) $1s^2 2s^2 2p^6 3s^2 3p^4$ _____

15) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ _____

16) $[Kr] 5s^2 4d^{10} 5p^3$ _____

16
S
sulfur
32.1

37
Rb
rubidium
85.5

51
Sb
antimony
121.8