

TOPICAL TEST 1 #01 - T14.2

Atomic structure and the Periodic table

1 Define the term 'atom'?

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2 State the subatomic particles in an atom.

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3 State the charges, relative masses the location of the subatomic particles in the atom.

Subatomic particles	charges	Relative mass	Location in the atom
PROTON			
			IN THE ELECTRON SHELL
	0		

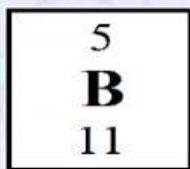
4 Define the term Proton number / atomic mass number

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.....

5 Define the term Nucleon number / mass number

.....
.....

6 What does this symbols and numbers represents?



5 =

B =

11 =

7 State the number of sub-particles in the following elements:

5 B 11	No. of Proton = No. of Electron = No. of Neutron =	35 Br 80	No. of Proton = No. of Electron = No. of Neutron =
17 Cl 35.5	No. of Proton = No. of Electron = No. of Neutron =	33 As 75	No. of Proton = No. of Electron = No. of Neutron =
19 K 39	No. of Proton = No. of Electron = No. of Neutron =	55 Cs 133	No. of Proton = No. of Electron = No. of Neutron =
13 Al 27	No. of Proton = No. of Electron = No. of Neutron =	12 Mg 24	No. of Proton = No. of Electron = No. of Neutron =
14 Si 28	No. of Proton = No. of Electron = No. of Neutron =	36 Kr 84	No. of Proton = No. of Electron = No. of Neutron =

8 Write the electronic structure for the following elements:

